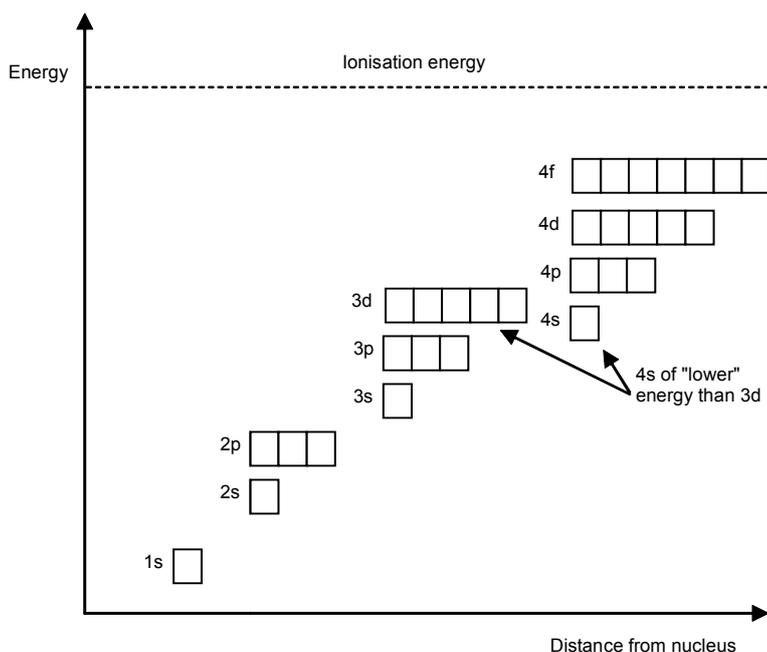




ELECTRON ARRANGEMENT 1

How electrons are arranged in atoms and ions

- Electrons are arranged in electron shells (energy levels), which themselves have sub-shells (sub-levels). The diagram below represents the first four main shells of electrons (the exact energies and distance from the nucleus vary and are simplified).



- Each sub-level consists of electron orbitals (region of space in which the electron spends most of its time).
- Each orbital can hold 2 electrons with opposite spins (one electron spins clockwise and one anticlockwise).
- Orbitals are regions of space that electrons are most likely to be in. The shapes of some orbitals are shown, but you do **not** have to be able to draw these shapes.

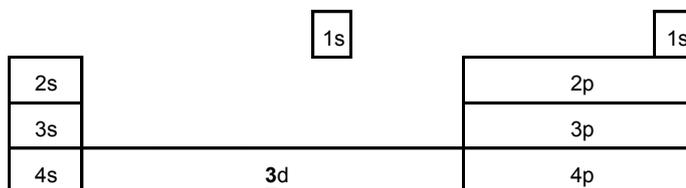
Sub-level	Number of orbitals in sub-level	Shape (no need to learn)	Maximum number of electrons in sub-level
s	1		2
p	3		6
d	5		10
f	7	Even more complicated!	14

- The way the electron orbitals are filled up:

1) Electrons enter the lowest energy orbital available (**Aufbau principle**).

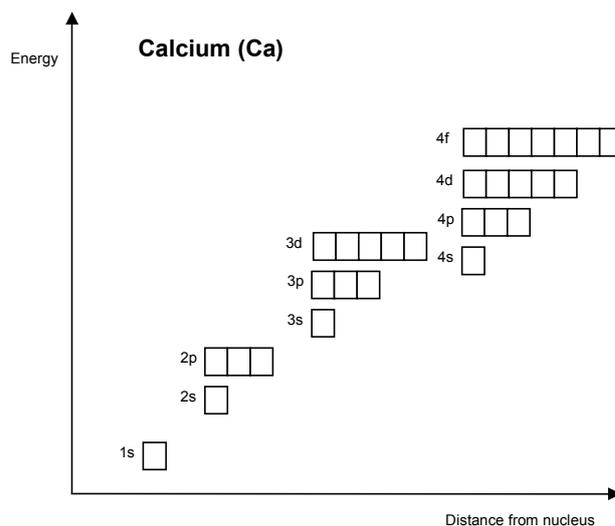
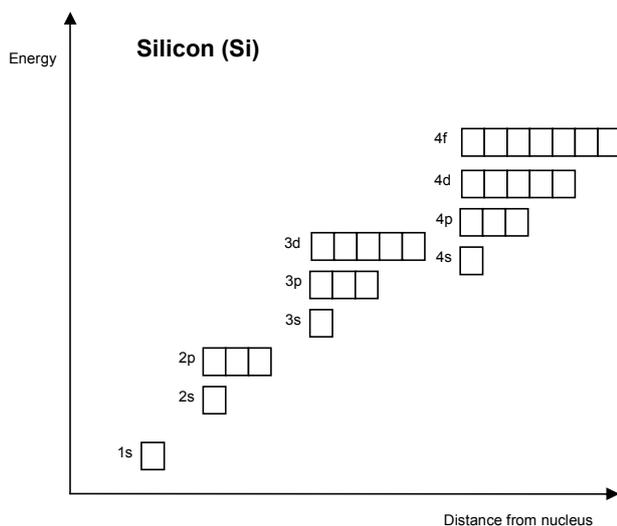
This diagram helps you to work out the order in which orbitals fill:
1s, 2s, 2p, 3s, 3p, 4s, 3d, 4p,

However, it can be easier to read across the periodic table, but remember that the first transition metal row is 3d:



2) Electrons prefer to occupy orbitals on their own, and only pair up when no empty orbitals of the same energy are available (**Hund's Rule**).

Complete the diagrams below to show the electron structure of the atoms shown:



- In ions, the electrons in the highest energy levels are lost first, but note that when losing electrons, electrons are lost from 4s before 3d (the energy levels are very close, and when electrons fill them, 4s goes above 3d).
- Note these two exceptions to the expected pattern, both of which stem from the 4s and 3d levels being very close in energy:



This is a slightly lower energy arrangement as the reduced e^-e^- repulsion makes up for the fact one electron is in a slightly higher energy level.



This is a slightly lower energy arrangement.

Complete the electron configurations **Electron Arrangement 2 (Chemsheets AS 1004)**