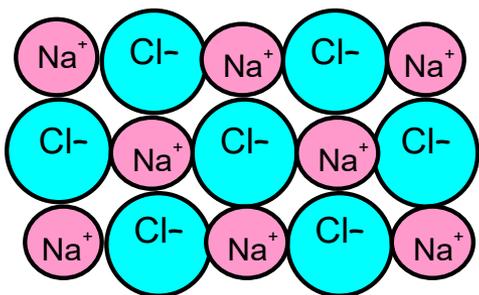




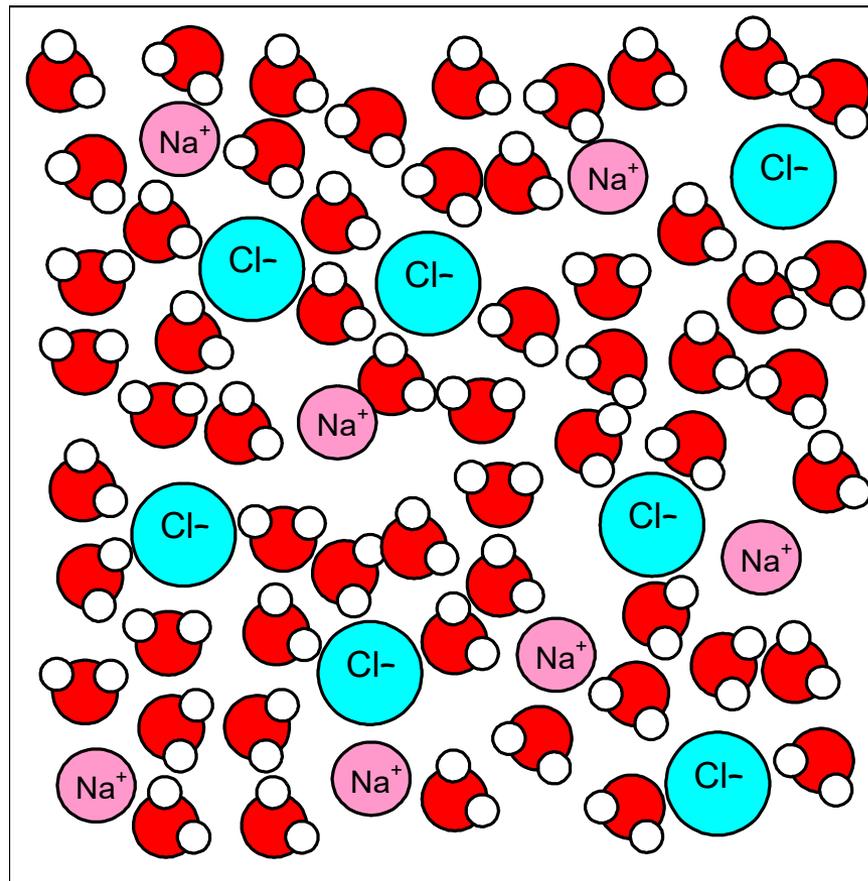
[WWW.CHEMSHEETS.CO.UK](http://www.chemsheets.co.uk)

IONIC EQUATIONS

When an IONIC substance dissolves, the ions separate and mix in with the water:

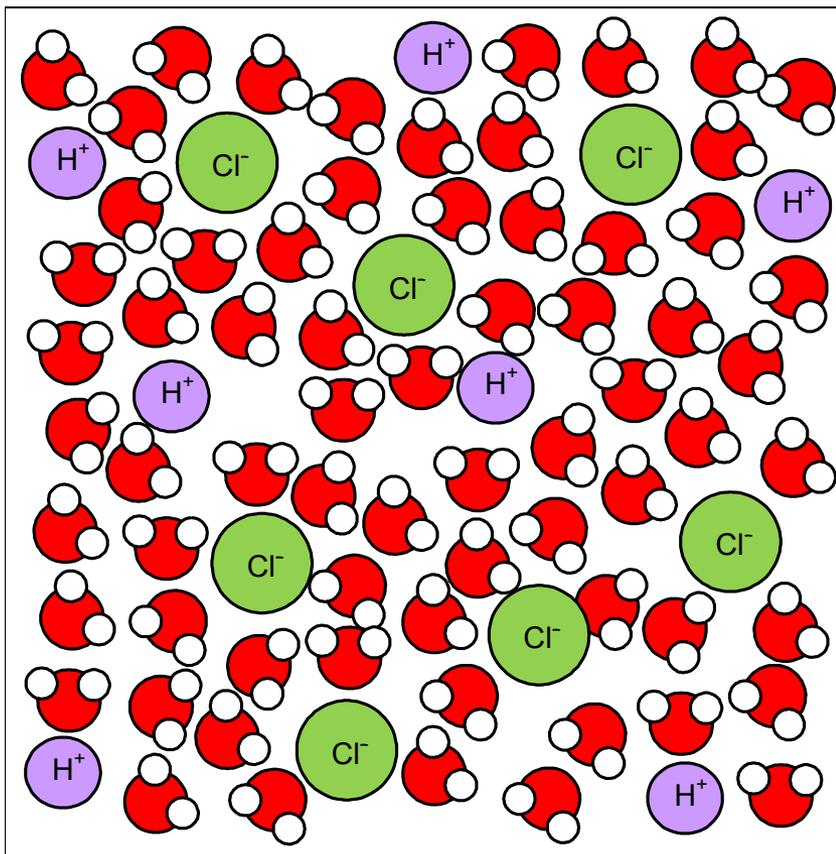


NaCl(s)

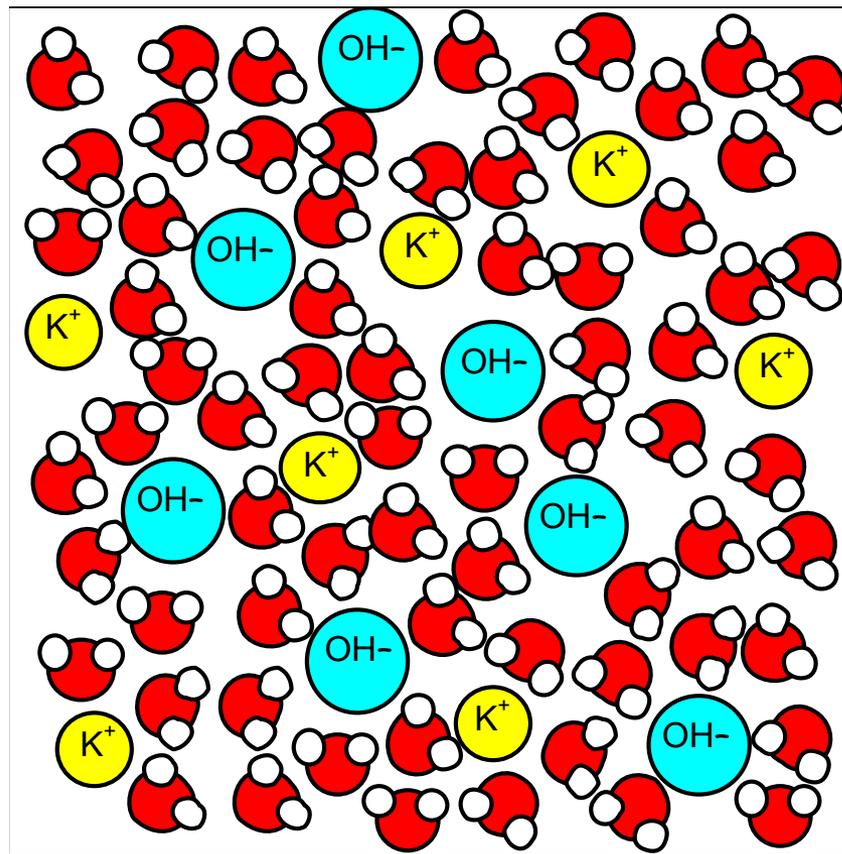


NaCl(aq)

ACID + ALKALI

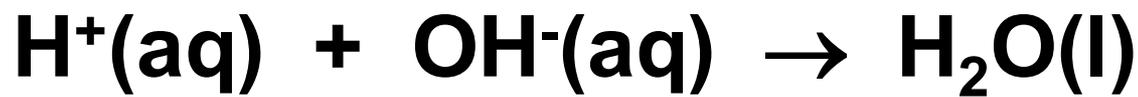
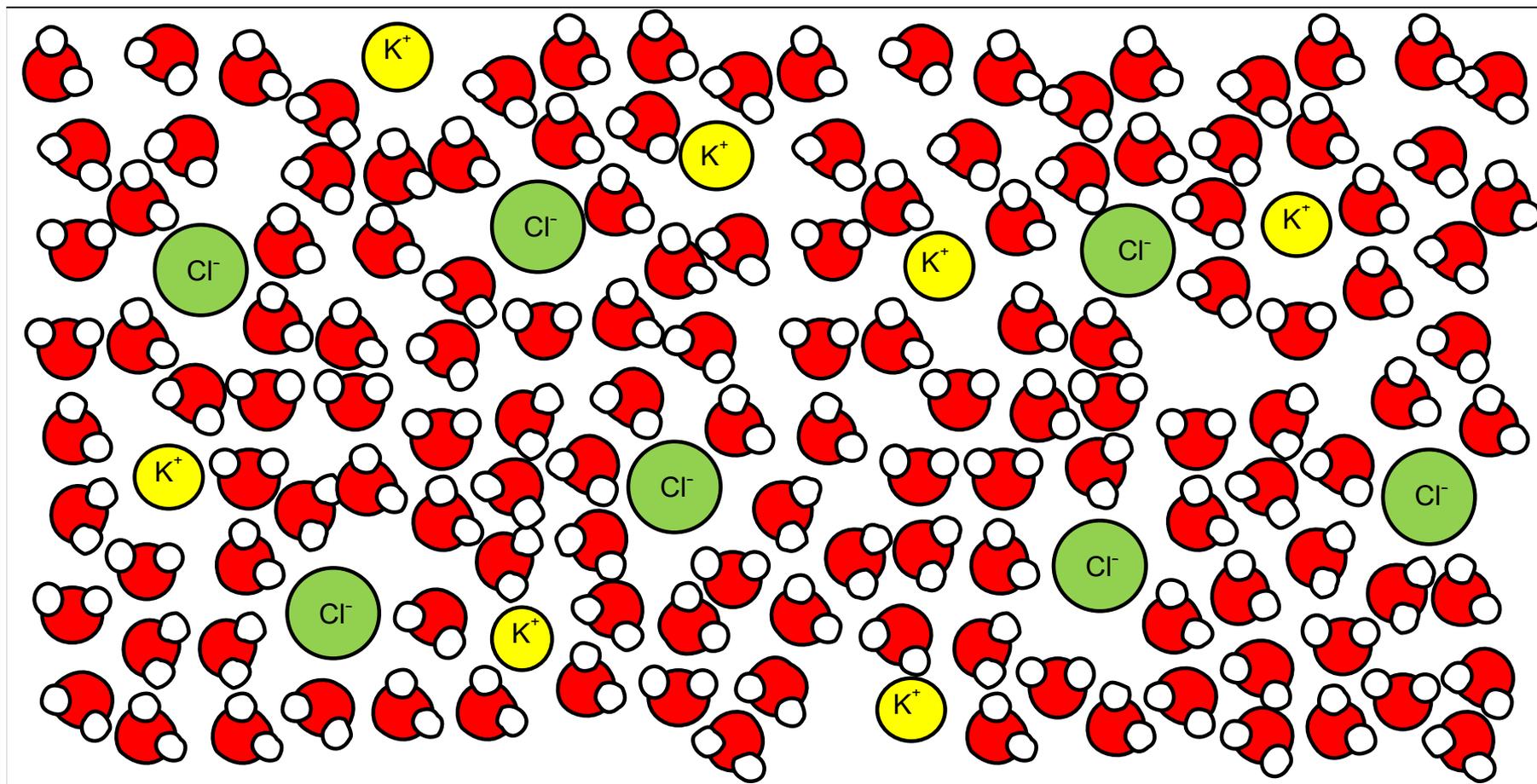


HCl(aq)

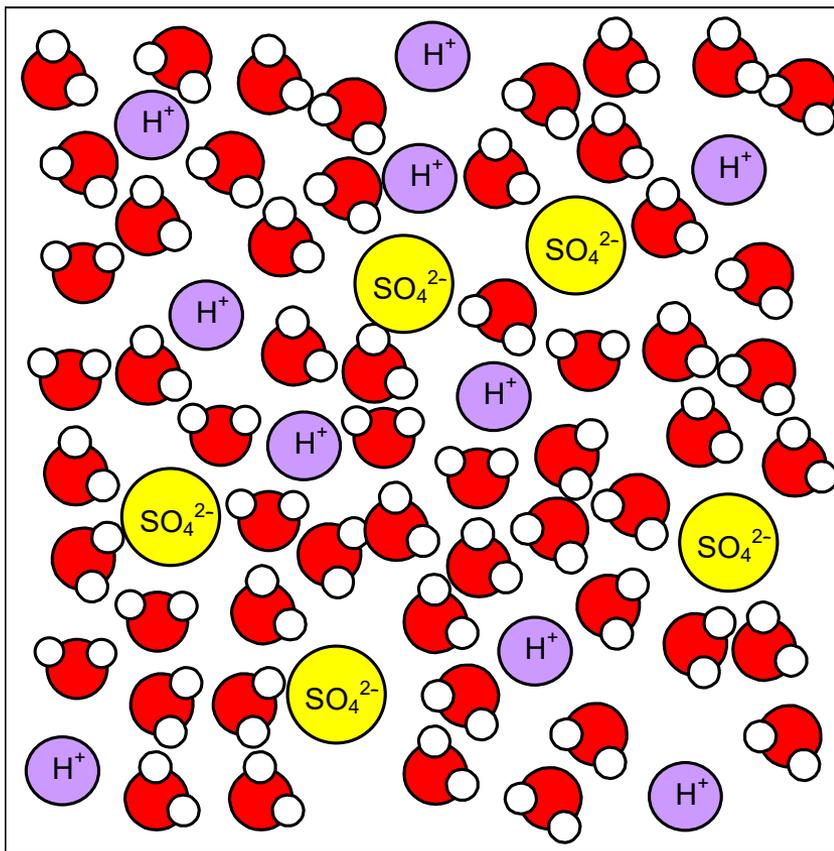


KOH(aq)

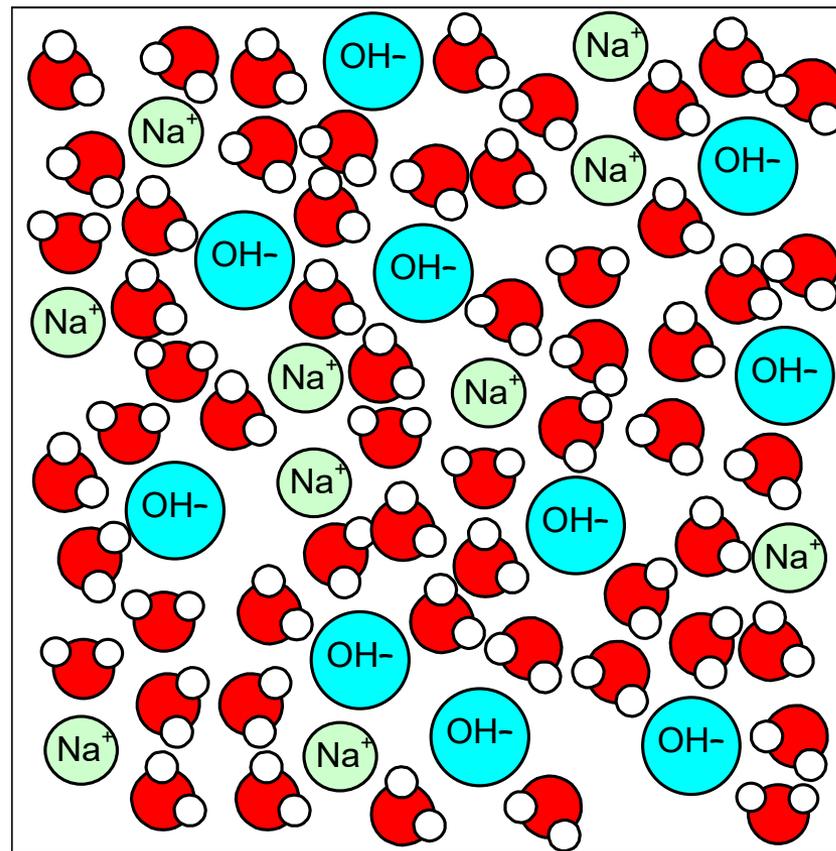
ACID + ALKALI



ACID + ALKALI

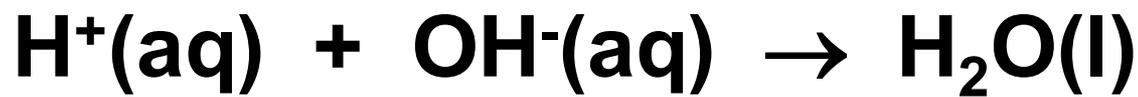
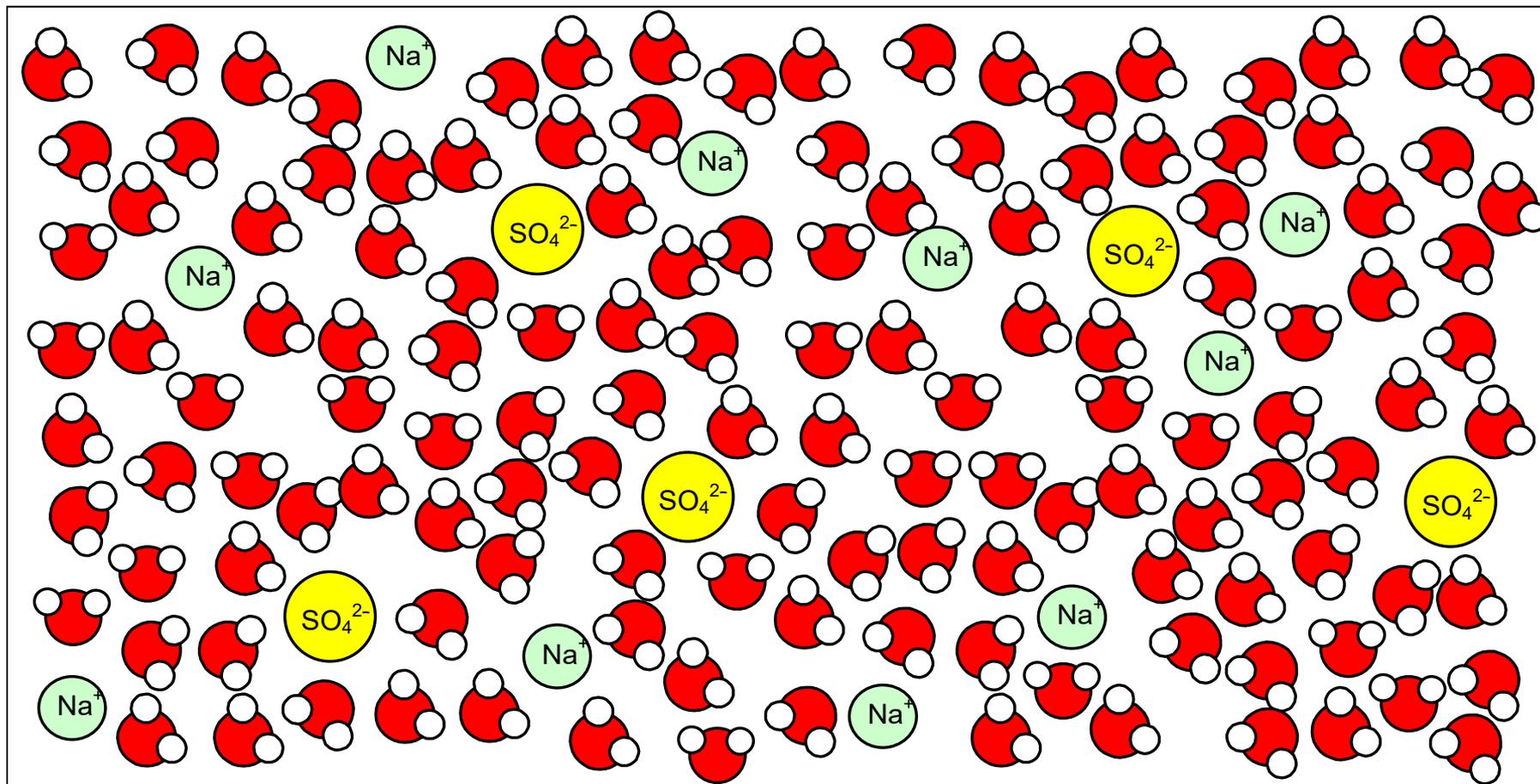


$\text{H}_2\text{SO}_4(\text{aq})$

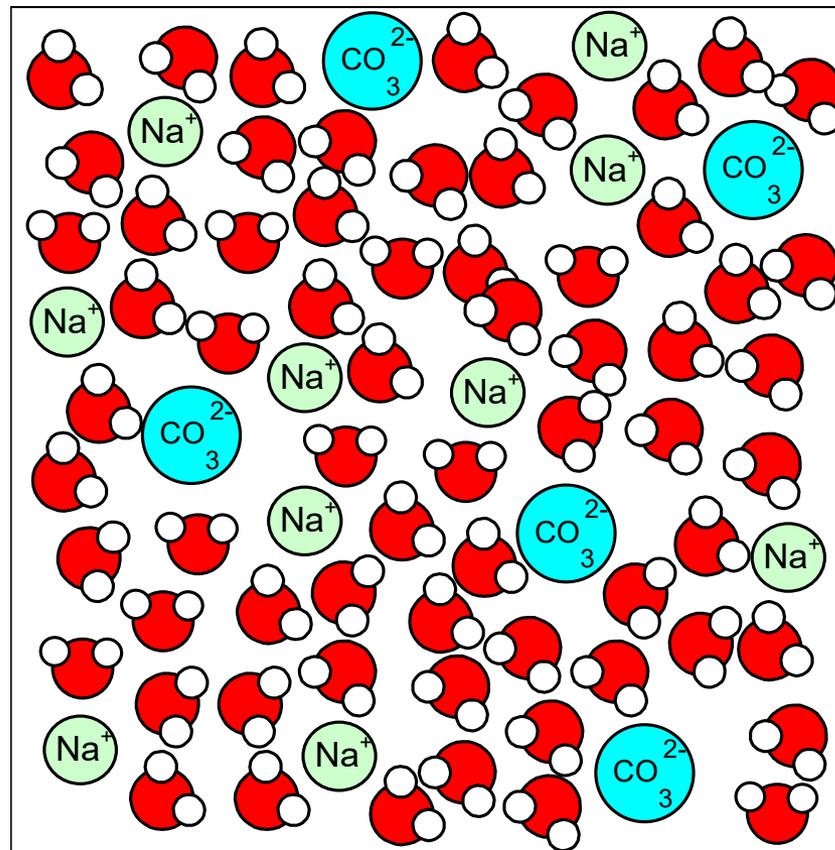
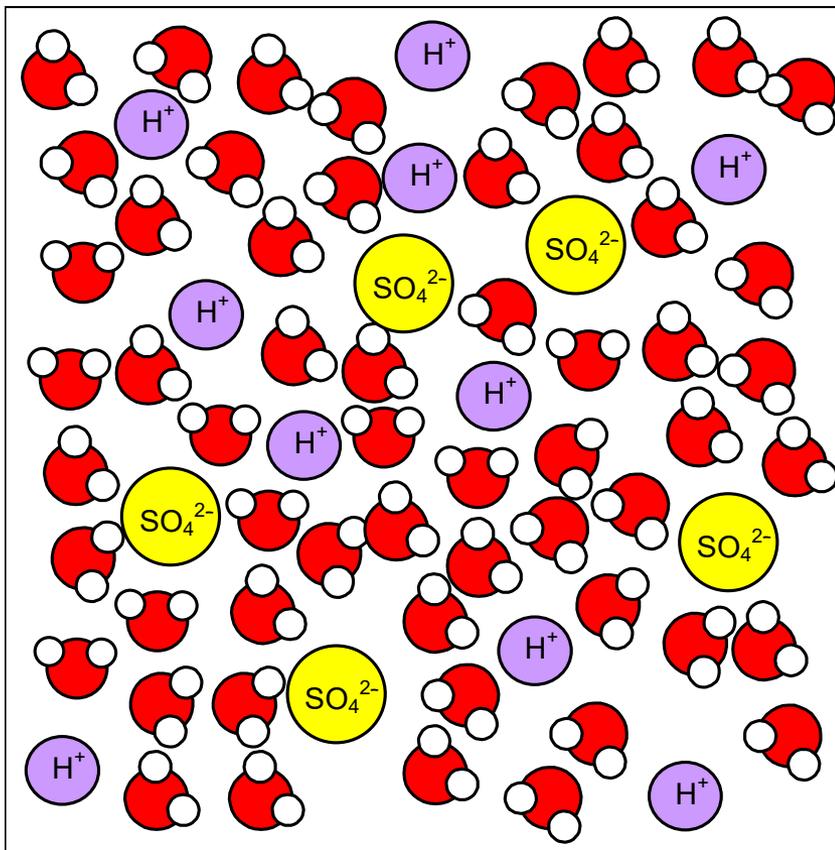


$\text{NaOH}(\text{aq})$

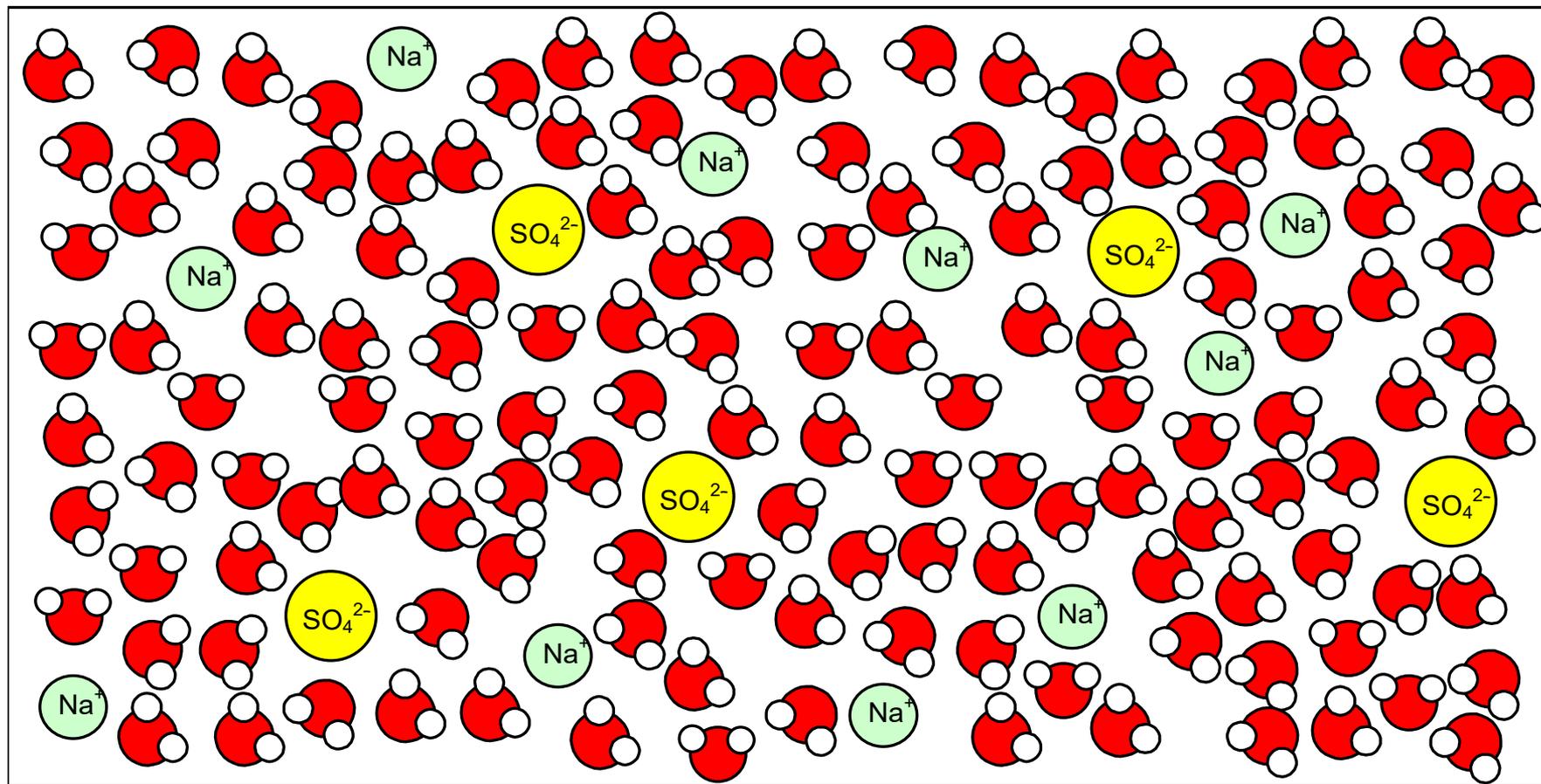
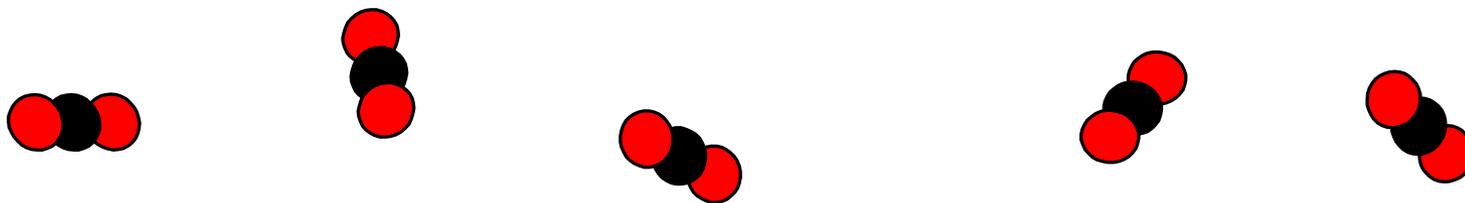
ACID + ALKALI



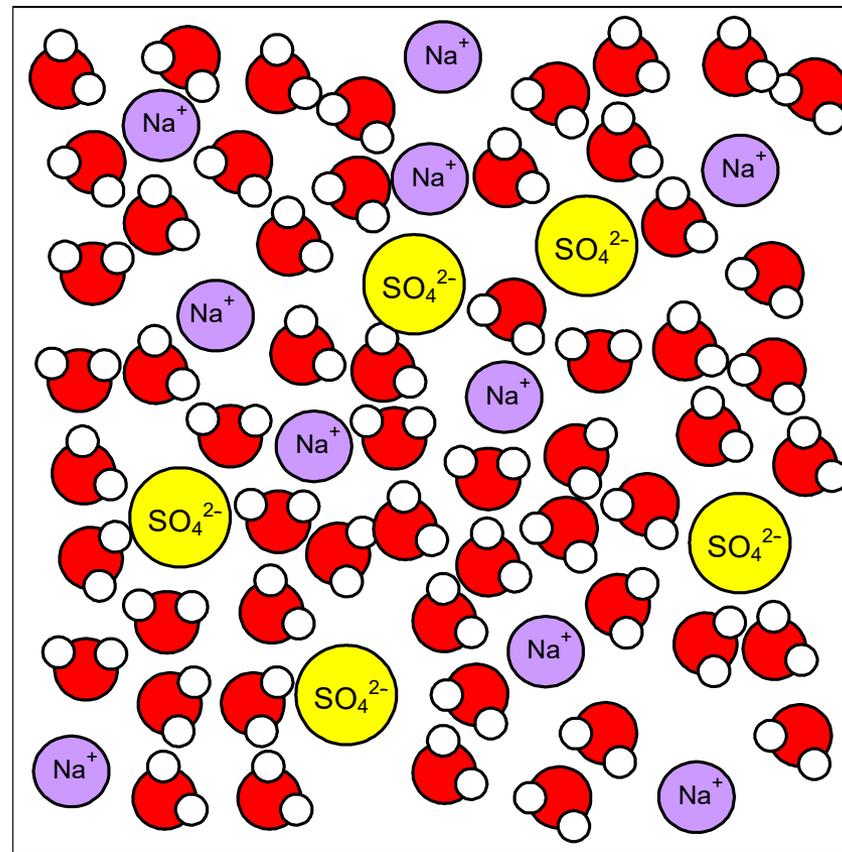
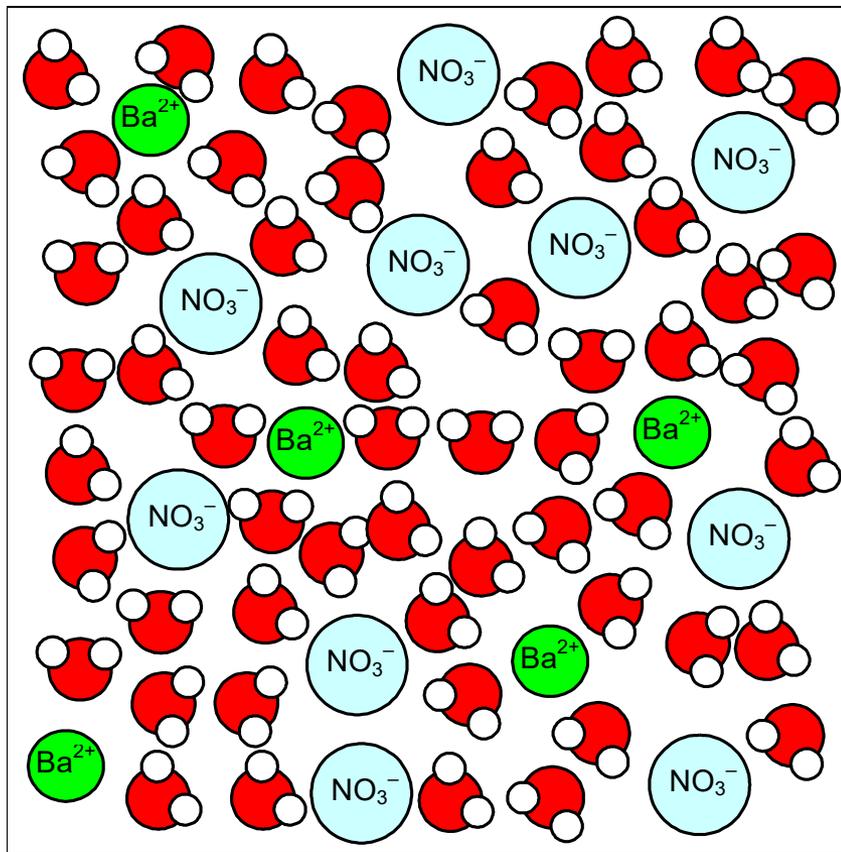
ACID + CARBONATE



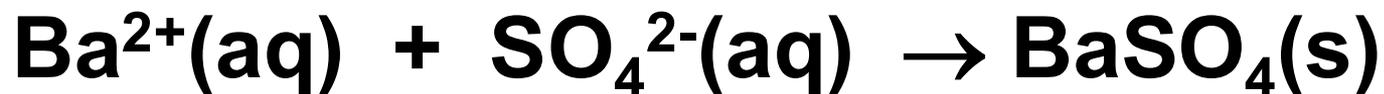
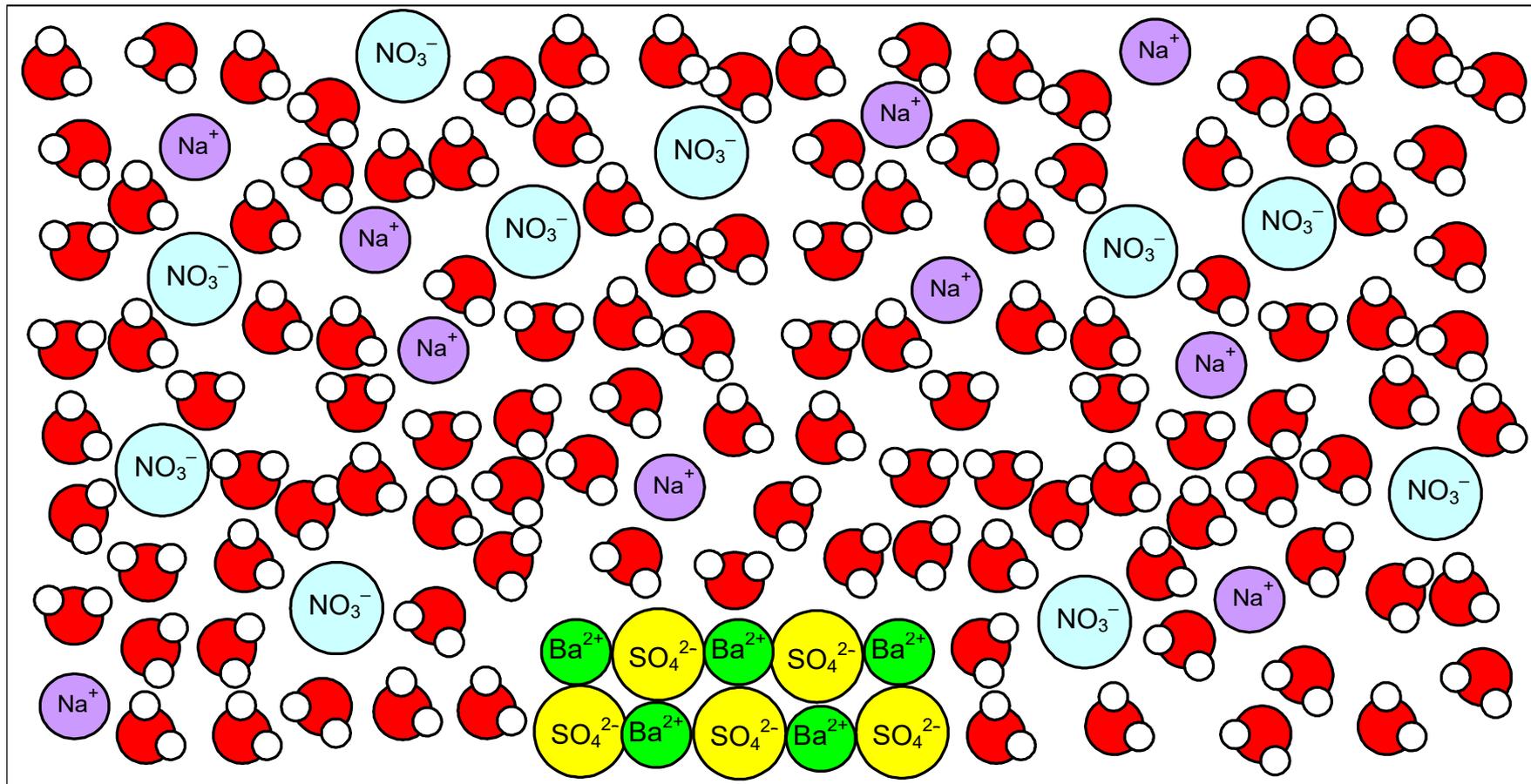
ACID + CARBONATE



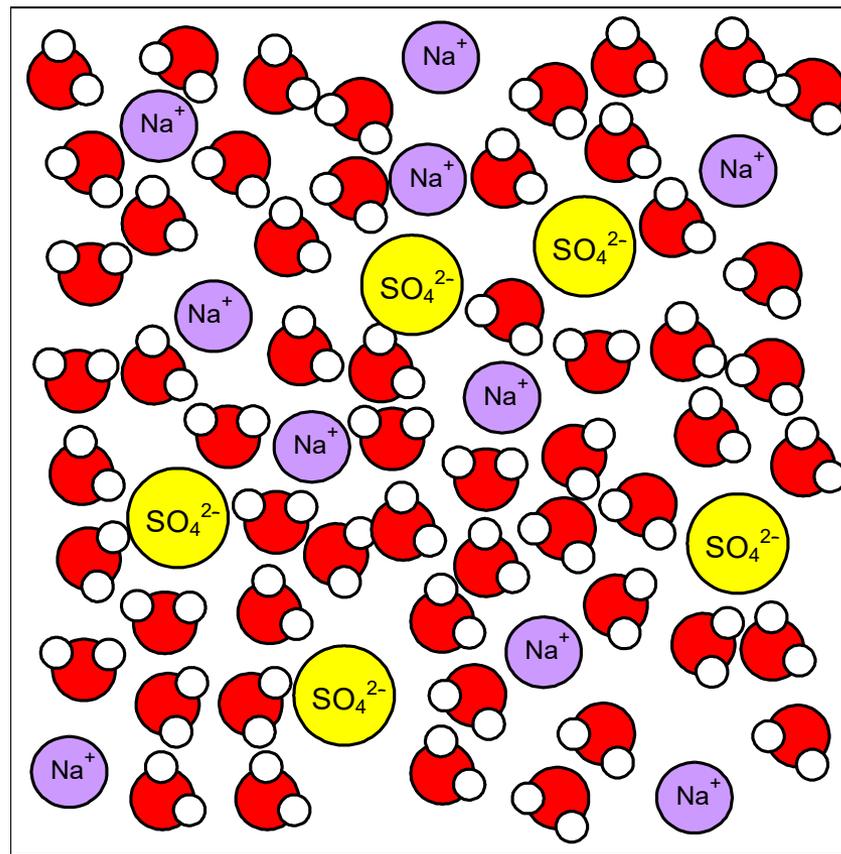
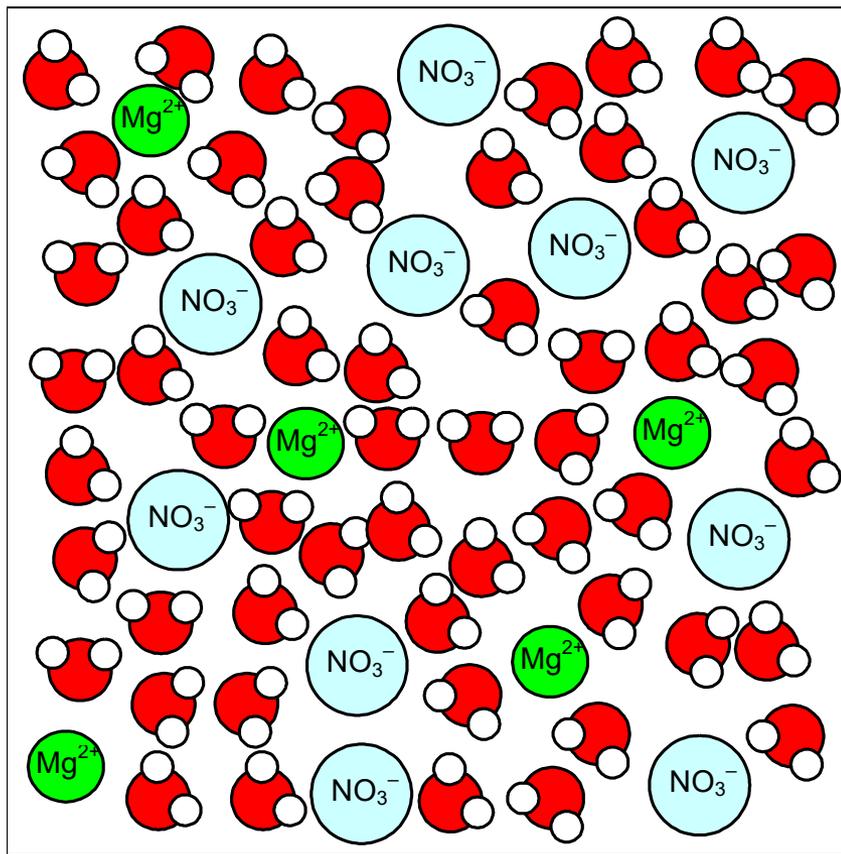
PRECIPITATION REACTION



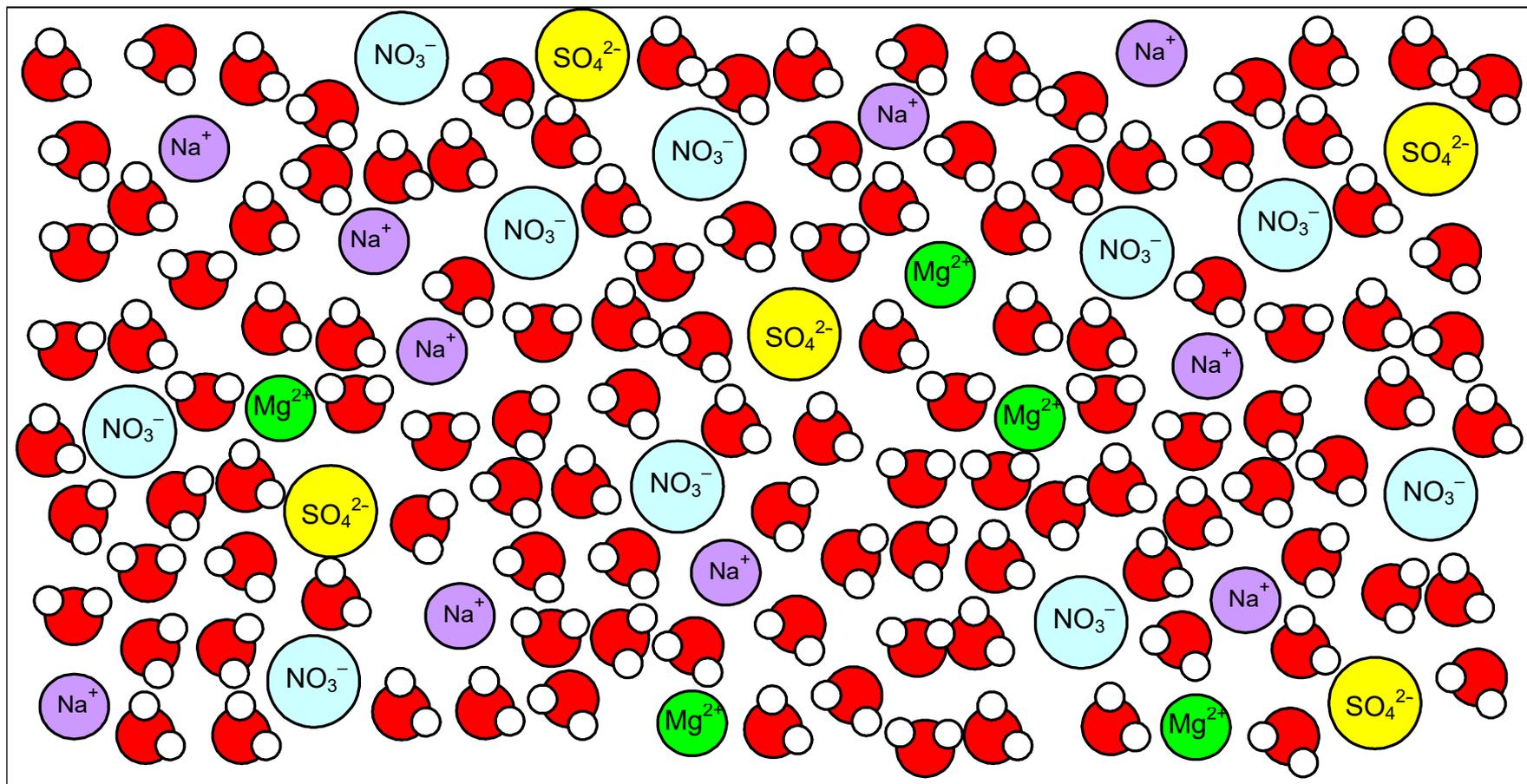
PRECIPITATION REACTION



PRECIPITATION REACTION



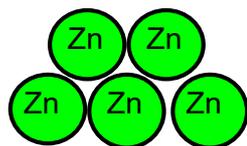
PRECIPITATION REACTION



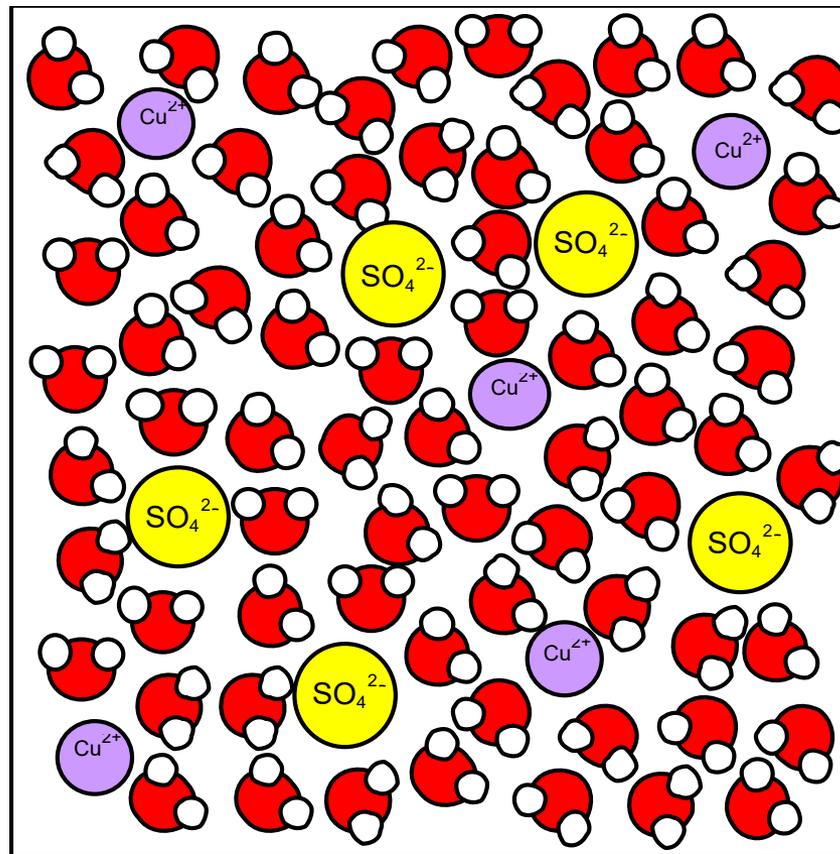
No reaction

(none of the ion combinations are insoluble)

"DISPLACEMENT" REACTION



Zn(s)



$\text{CuSO}_4(\text{aq})$

"DISPLACEMENT" REACTION

