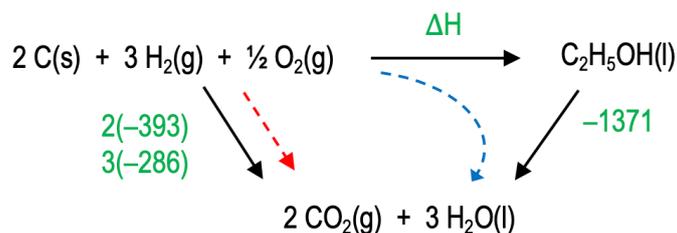




HESS'S LAW 2 - COMBUSTION

e.g. 1

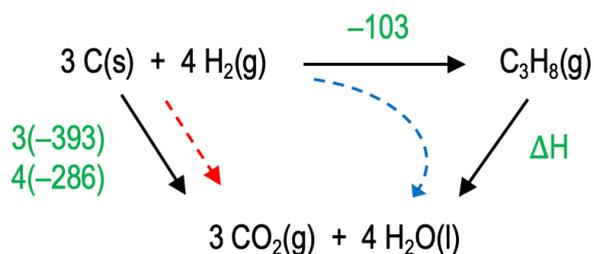


$$\Delta\text{H} - 1371 = 2(-393) + 3(-286)$$

$$\Delta\text{H} = 1371 + 2(-393) + 3(-286)$$

$$= -273 \text{ kJ mol}^{-1}$$

e.g. 2

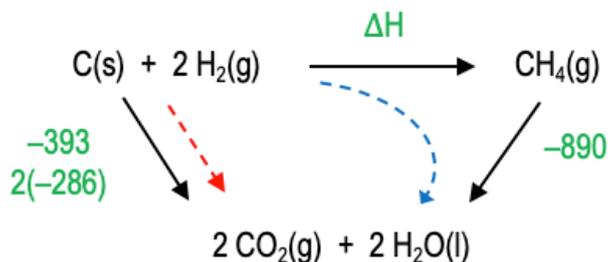


$$-103 + \Delta\text{H} = 3(-393) + 4(-286)$$

$$\Delta\text{H} = 103 + 3(-393) + 4(-286)$$

$$= -2220 \text{ kJ mol}^{-1}$$

1

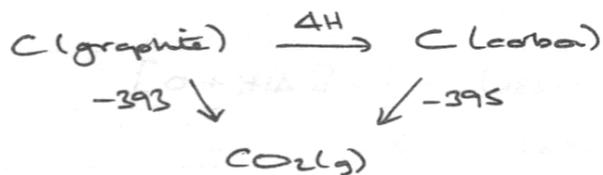


$$-890 + \Delta\text{H} = -393 + 2(-286)$$

$$\Delta\text{H} = 890 - 393 + 2(-286)$$

$$= -75 \text{ kJ mol}^{-1}$$

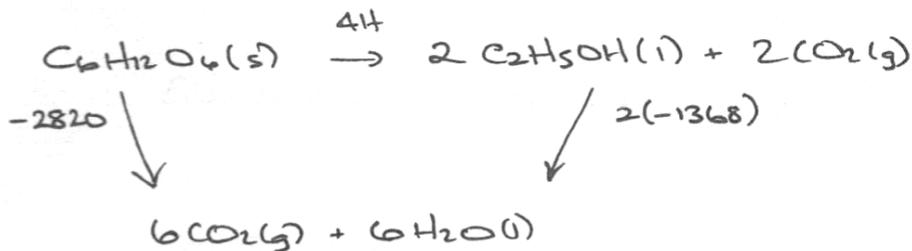
2



$$\Delta H - 395 = -393$$

$$\Delta H = \underline{+2 \text{ kJ/mol}}$$

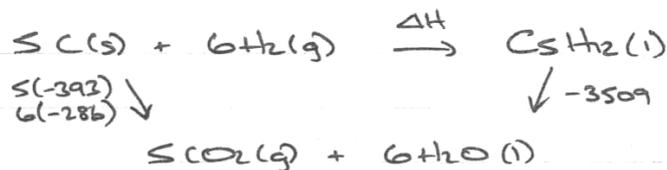
3



$$\Delta H + 2(-1368) = -2820$$

$$\Delta H = \underline{-84 \text{ kJ/mol}}$$

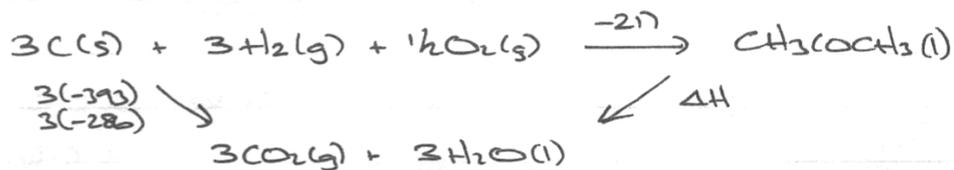
4



$$\Delta H - 3509 = 5(-393) + 6(-286)$$

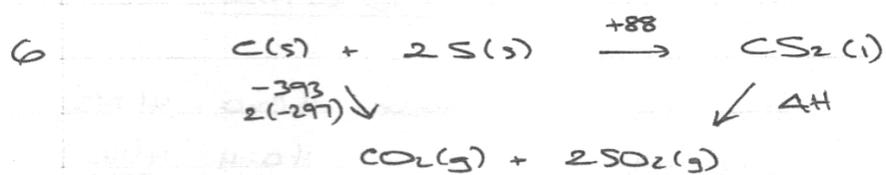
$$\Delta H = \underline{-172 \text{ kJ/mol}}$$

5



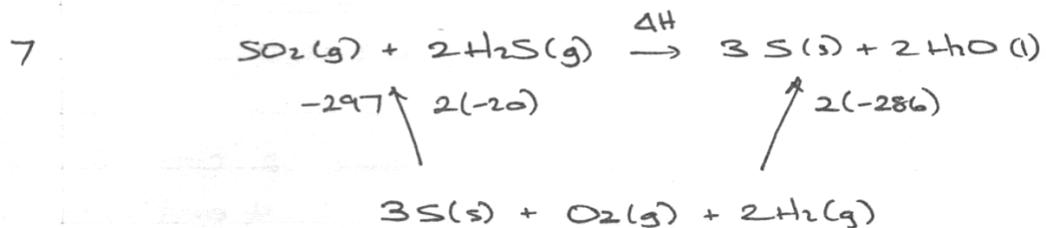
$$\Delta H - 217 = 3(-393) + 3(-286)$$

$$\Delta H = \underline{-1820 \text{ kJ/mol}}$$



$$\Delta H + 88 = -393 + 2(-297)$$

$$\Delta H = \underline{-1075 \text{ kJ/mol}}$$



$$\Delta H - 297 + 2(-20) = 2(-286)$$

$$\Delta H = \underline{-235 \text{ kJ/mol}}$$