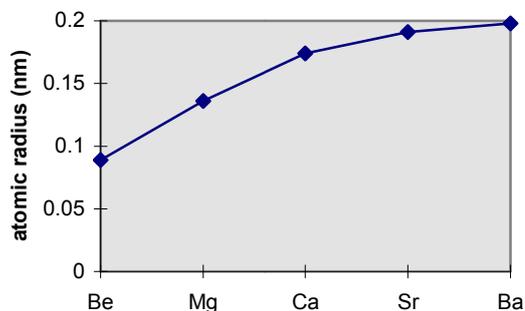




GROUP 2 – PHYSICAL PROPERTIES

Atomic radius

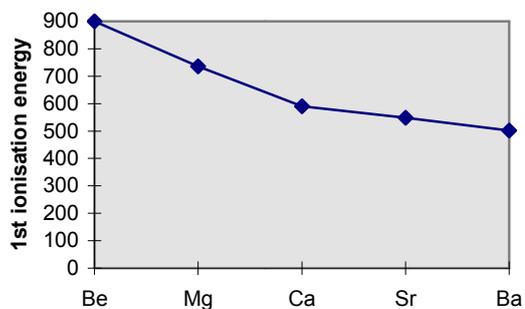
- atomic radius gets bigger down the group
- more shells of electrons



1st ionisation energy

Enthalpy change to remove one electron from each atom in a mole of gaseous atoms.

- atomic radius gets bigger down the group
- more shielding down the group
- attraction between nucleus and outer shell electron decreases down the group
- therefore lower ionisation energy down the group



Melting points

- lower melting points down the group as weaker metallic bonding
- weaker metallic bonding as atoms (ions) get bigger
- change in trend after Mg due to change in way atoms (ions) pack together (pack together in a closer arrangement)

