



[WWW.CHEMSHEETS.CO.UK](http://www.chemsheets.co.uk)

# HALIDE ION TESTS

Halide ion	Action of $\text{AgNO}_3(\text{aq})$	Action of dilute $\text{NH}_3(\text{aq})$	Action of conc. $\text{NH}_3(\text{aq})$
$\text{F}^-(\text{aq})$			
$\text{Cl}^-(\text{aq})$			
$\text{Br}^-(\text{aq})$			
$\text{I}^-(\text{aq})$			

Halide ion	Action of $\text{AgNO}_3(\text{aq})$	Action of dilute $\text{NH}_3(\text{aq})$	Action of conc. $\text{NH}_3(\text{aq})$
$\text{F}^-(\text{aq})$	No precipitate		
$\text{Cl}^-(\text{aq})$			
$\text{Br}^-(\text{aq})$			
$\text{I}^-(\text{aq})$			

Halide ion	Action of AgNO <sub>3</sub> (aq)	Action of dilute NH <sub>3</sub> (aq)	Action of conc. NH <sub>3</sub> (aq)
F <sup>-</sup> (aq)	No precipitate		
Cl <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ white ppt		
Br <sup>-</sup> (aq)			
I <sup>-</sup> (aq)			

Halide ion	Action of AgNO <sub>3</sub> (aq)	Action of dilute NH <sub>3</sub> (aq)	Action of conc. NH <sub>3</sub> (aq)
F <sup>-</sup> (aq)	No precipitate		
Cl <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ white ppt	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow$ $[\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$	
Br <sup>-</sup> (aq)			
I <sup>-</sup> (aq)			

Halide ion	Action of AgNO <sub>3</sub> (aq)	Action of dilute NH <sub>3</sub> (aq)	Action of conc. NH <sub>3</sub> (aq)
F <sup>-</sup> (aq)	No precipitate		
Cl <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ white ppt	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$
Br <sup>-</sup> (aq)			
I <sup>-</sup> (aq)			

Halide ion	Action of AgNO <sub>3</sub> (aq)	Action of dilute NH <sub>3</sub> (aq)	Action of conc. NH <sub>3</sub> (aq)
F <sup>-</sup> (aq)	No precipitate		
Cl <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ white ppt	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$
Br <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq}) \rightarrow \text{AgBr}(\text{s})$ cream ppt		
I <sup>-</sup> (aq)			

Halide ion	Action of AgNO <sub>3</sub> (aq)	Action of dilute NH <sub>3</sub> (aq)	Action of conc. NH <sub>3</sub> (aq)
F <sup>-</sup> (aq)	No precipitate		
Cl <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ white ppt	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$
Br <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq}) \rightarrow \text{AgBr}(\text{s})$ cream ppt	Insoluble	
I <sup>-</sup> (aq)			

Halide ion	Action of AgNO <sub>3</sub> (aq)	Action of dilute NH <sub>3</sub> (aq)	Action of conc. NH <sub>3</sub> (aq)
F <sup>-</sup> (aq)	No precipitate		
Cl <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ white ppt	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$
Br <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq}) \rightarrow \text{AgBr}(\text{s})$ cream ppt	Insoluble	Soluble $\text{AgBr}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Br}^-(\text{aq})$
I <sup>-</sup> (aq)			

Halide ion	Action of AgNO <sub>3</sub> (aq)	Action of dilute NH <sub>3</sub> (aq)	Action of conc. NH <sub>3</sub> (aq)
F <sup>-</sup> (aq)	No precipitate		
Cl <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ white ppt	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$
Br <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq}) \rightarrow \text{AgBr}(\text{s})$ cream ppt	Insoluble	Soluble $\text{AgBr}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Br}^-(\text{aq})$
I <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$ yellow ppt		

Halide ion	Action of AgNO <sub>3</sub> (aq)	Action of dilute NH <sub>3</sub> (aq)	Action of conc. NH <sub>3</sub> (aq)
F <sup>-</sup> (aq)	No precipitate		
Cl <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ white ppt	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$
Br <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq}) \rightarrow \text{AgBr}(\text{s})$ cream ppt	Insoluble	Soluble $\text{AgBr}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Br}^-(\text{aq})$
I <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$ yellow ppt	Insoluble	

Halide ion	Action of AgNO <sub>3</sub> (aq)	Action of dilute NH <sub>3</sub> (aq)	Action of conc. NH <sub>3</sub> (aq)
F <sup>-</sup> (aq)	No precipitate		
Cl <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ white ppt	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$	Soluble $\text{AgCl}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Cl}^-(\text{aq})$
Br <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq}) \rightarrow \text{AgBr}(\text{s})$ cream ppt	Insoluble	Soluble $\text{AgBr}(\text{s}) + 2 \text{NH}_3(\text{aq}) \rightarrow [\text{Ag}(\text{NH}_3)_2]^+(\text{aq}) + \text{Br}^-(\text{aq})$
I <sup>-</sup> (aq)	$\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$ yellow ppt	Insoluble	Insoluble

