



# OXIDISING POWER OF HALOGENS

In each reaction, place about 2 cm depth of the halide solution in a test tube and then add the halogen solution dropwise. Look carefully for colour changes, bearing in mind that if no reaction takes place the halogen solution will have been diluted and become paler in colour. **Dispose of solutions down the fume cupboard sink.**



Chlorine  
Bromine



Bromine



Iodine

Halide (aq)	Halogen (aq)			
	Cl <sub>2</sub> (aq)	Br <sub>2</sub> (aq)	I <sub>2</sub> (aq)	
Cl <sup>-</sup> (aq)				Observation
				Ionic equation
Br <sup>-</sup> (aq)				Observation
				Ionic equation
I <sup>-</sup> (aq)				Observation
				Ionic equation

## Oxidising power:

## Explanation

Strongest .....

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Weakest .....

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.....

.....

Cl

Br

I