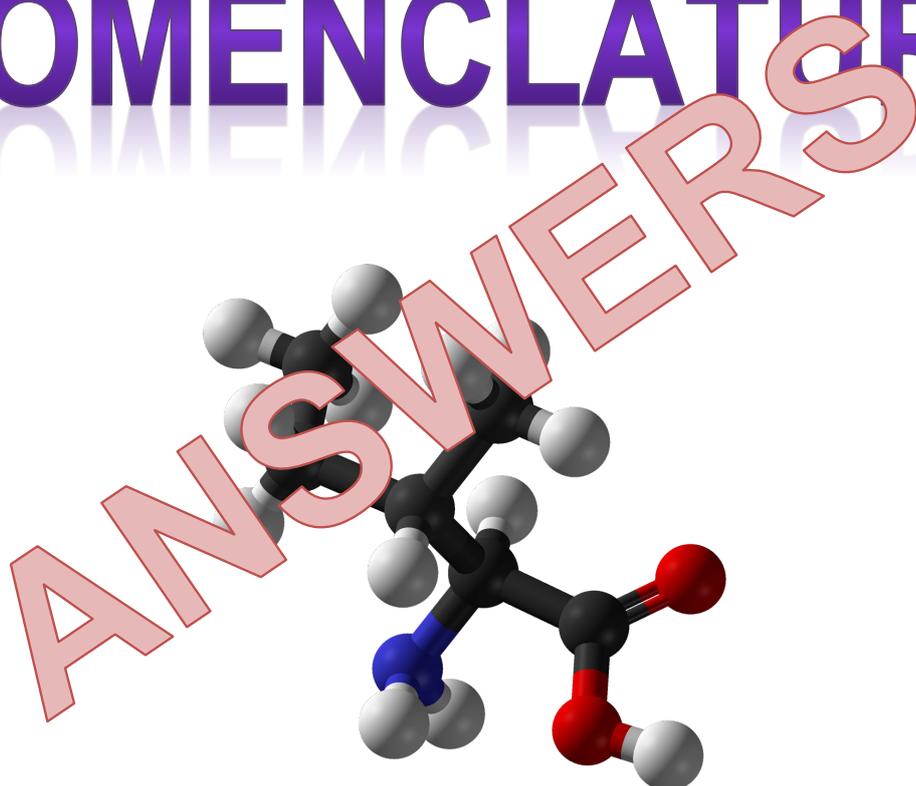
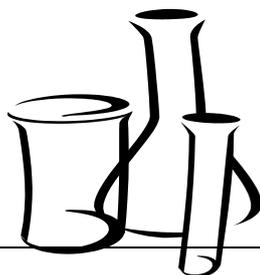


ORGANIC NOMENCLATURE



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The naming of organic compounds is controlled by IUPAC (International Union of Pure and Applied Chemistry). This gives a very systematic way of naming all organic molecules.

Some guidance on naming organic molecules

1 Longest C chain

- look at the molecule to find the longest C chain – if there is a functional group it must be the longest C chain containing the functional group
- note that alkyl groups (e.g. methyl, ethyl, propyl) are not functional groups.

number of C atoms	1C	2C	3C	4C	5C	6C
start of name	meth	eth	prop	but	pent	hex

2 Number the C chain

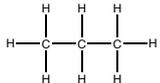
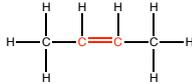
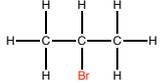
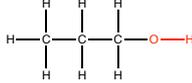
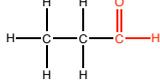
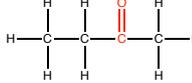
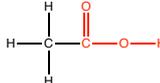
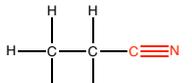
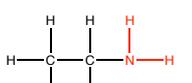
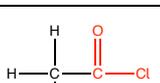
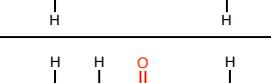
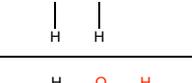
- number the longest C chain (that contains the functional group) from the end that gives the functional group(s) the lowest numbers
- note that some functional groups (labelled * in the chart, e.g. aldehyde group) are carbon-1 in the chain by definition (and do not need a number themselves therefore), and define which end the numbers start for other groups
- note that we should only include numbers in names when they are necessary – but we usually only omit numbers in the simplest molecules where it is obvious that no number is needed

3 Putting the groups on the chain on the name

- groups that come off the C chain are given a number to show which C atom they are on (if necessary)
- when we write the name, these groups are listed in alphabetical order
- if there are several groups that are the same, then the prefix di (for two the same), tri (for three the same), tetra (for four the same), etc are added first (note that the prefix di/tri/tetra/etc is ignored when we work out alphabetical order)

Note

- numbers are separated by commas
- numbers and letters are separated by dashes

functional group	prefix	suffix	example name
alkane		ane	 propane
alkene		ene	 but-2-ene
halogenoalkane	fluoro / chloro / bromo / iodo		 2-bromopropane
alcohol	hydroxy	ol	 propan-1-ol
aldehyde *		al	 propanal
ketone	oxo	one	 butanone
carboxylic acid *		oic acid	 ethanoic acid
nitrile *		nitrile	 propanenitrile
amine	amino	amine	 ethanamine (ethylamine)
acyl chloride *		oyl chloride	 ethanoyl chloride
acid anhydride		oic anhydride	 ethanoic anhydride
ester		oate	 methyl propanoate
amide		amide	 ethanamide

* indicates this the C atom in this functional group is defined as carbon-1 in the chain

shading indicates that this prefix is only used if the suffix cannot be used

	Structural formula	Skeletal formula	Name & comments
5	$ \begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3 - \text{CH} - \text{C} - \text{CH}_2 - \text{CH}_3 \\ \quad \\ \text{CH}_3 \quad \text{CH}_2 \\ \\ \text{CH}_3 \end{array} $		<p>Longest C chain is 5C atoms – so it is a <i>pentane</i></p> <p>There are two methyl groups and one ethyl group coming off the chain.</p> <p>If the chain is counted from the left end, then we have <i>2-methyl</i>, <i>3-ethyl</i> and <i>3-methyl</i>. If it is numbered from the right we have <i>3-methyl</i>, <i>3-ethyl</i> and <i>4-methyl</i>. We get lower numbers if we number from the left end.</p> <p>The groups are listed in alphabetical order, but we ignore the <i>di</i> of the dimethyl. Therefore, <i>ethyl</i> comes before <i>dimethyl</i> alphabetically.</p> <p>3-ethyl-2,3-dimethylpentane</p>
	Displayed formula		

	Structural formula	Skeletal formula	Name & comments
6	$ \begin{array}{c} \text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH} - \text{CH}_3 \\ \\ \text{CH}_2 \\ \\ \text{CH}_3 \end{array} $		<p>Longest C chain is 6C atoms (not 5) – so it is a <i>hexane</i></p> <p>There is one methyl group coming off the chain.</p> <p>If the chain is counted from the left end, then we have <i>4-methyl</i>. If it is numbered from the bottom we have <i>3-methyl</i>. We get the lower number if we number from the bottom.</p> <p>3-methylhexane</p>
	Displayed formula		

	Structural formula	Skeletal formula	Displayed formula
7	$ \begin{array}{c} \text{CH}_3 \quad \text{CH}_3 \quad \text{CH}_3 \\ \quad \quad \\ \text{CH}_3 - \text{CH}_2 - \text{CH} - \text{CH} - \text{CH} - \text{CH}_3 \end{array} $		
	Name 2,3,4-trimethylhexane		

	Structural formula	Skeletal formula	Displayed formula
8	$ \begin{array}{c} \text{CH}_3 \quad \quad \text{CH}_3 \\ \quad \quad \\ \text{CH}_3 - \text{CH} - \text{CH}_2 - \text{C} - \text{CH}_3 \\ \\ \text{CH}_2 \\ \\ \text{CH}_3 \end{array} $		
	Name 2,4,4-trimethylhexane		

ALKENES

	Structural formula	Skeletal formula	Name & comments
9	$\text{CH}_3-\text{CH}_2-\text{CH}=\text{CH}_2$		Longest C chain that contains the C=C functional group is 4C atoms – so it has a <i>but</i> stem. It has an alkene functional group and so is a <i>butene</i> .
	Displayed formula		The C=C group could be in different places so needs a number. If we count from the right it is between carbon-1 and carbon-2. If we count from the left it is between carbon-3 and carbon-4. We count from the right therefore to get the lower number, so it is given number 1. but-1-ene
	$\begin{array}{cccc} & \text{H} & \text{H} & \text{H} & \text{H} \\ & & & & \\ \text{H} & -\text{C} & -\text{C} & -\text{C} & =\text{C} \\ & & & & \\ & \text{H} & \text{H} & & \text{H} \end{array}$		

	Structural formula	Skeletal formula	Name & comments
10	$\begin{array}{ccccccc} & & \text{CH}_3 & & \text{CH}_3 & & \\ & & & & & & \\ \text{CH}_3 & -\text{CH} & -\text{C} & =\text{CH} & -\text{CH}_3 \end{array}$		Longest C chain that contains the C=C functional group is 5C atoms – so it has a <i>pent</i> stem. It has an alkene functional group and so is a <i>pentene</i> .
	Displayed formula		We have two methyl groups and a C=C double bond that need numbers. The methyl groups are not functional groups, so the only functional group is the C=C. We count from the right to give the C=C the lowest number. This means that we have <i>-2-ene</i> , <i>3-methyl</i> and <i>4-methyl</i> . 3,4-dimethylpent-2-ene
	$\begin{array}{ccccccc} & & & \text{H} & \text{H} & & \\ & & & & & & \\ & & & \text{C} & & & \\ & & & & & & \\ \text{H} & -\text{C} & -\text{C} & -\text{C} & =\text{C} & -\text{C} & -\text{H} \\ & & & & & & \\ & \text{H} & \text{H} & & & \text{H} & \\ & & & \text{C} & & & \\ & & & & & & \\ & & & \text{H} & \text{H} & & \end{array}$		

	Structural formula	Skeletal formula	Name & comments
11	$\begin{array}{ccccccc} & & \text{CH}_3 & & & \text{CH}_2 & \\ & & & & & & \\ \text{CH}_3 & -\text{CH} & -\text{CH}_2 & -\text{CH}_2 & -\text{C} & -\text{CH}_2 & -\text{CH}_3 \end{array}$		The longest C chain here has 7 C atoms, BUT it does not contain any functional group. The longest C chain that does contain the C=C functional group is 6 C atoms and so is a <i>hexene</i> .
	Displayed formula		To get the lowest number for the C=C functional group, the numbering starts from the end with the C=C, and so it is <i>-1-ene</i> . There are a methyl group on carbon-5 and so it is <i>5-methyl</i> . There is also an ethyl group on carbon-2 and so it is <i>2-ethyl</i> . <i>2-ethyl</i> comes before <i>5-methyl</i> due to alphabetical order. 2-ethyl-5-methylhex-1-ene
	$\begin{array}{ccccccc} & & \text{H} & & \text{H} & & \\ & & & & & & \\ & & \text{C} & & \text{C} & & \\ & & & & & & \\ \text{H} & -\text{C} & -\text{C} & -\text{C} & -\text{C} & =\text{C} & -\text{C} & -\text{H} \\ & & & & & & & \\ & \text{H} & \text{H} & \text{H} & \text{H} & & \text{H} & \end{array}$		

	Structural formula	Skeletal formula	Displayed formula
12	$\text{CH}_2=\text{CH}-\text{CH}_3$		$\begin{array}{ccc} \text{H} & \text{H} & \text{H} \\ & & \\ \text{C} & =\text{C} & -\text{C} & -\text{H} \\ & & \\ \text{H} & & \text{H} \end{array}$
	Name propene		

	Structural formula	Skeletal formula	Displayed formula
13	$\begin{array}{ccc} & \text{CH}_3 & \\ & & \\ \text{CH}_3 & -\text{C} & =\text{CH} & -\text{CH}_3 \end{array}$		$\begin{array}{cccc} & & \text{H} & \text{H} & \text{H} \\ & & & & \\ & & \text{C} & & \text{C} & & \text{C} \\ & & & & & & \\ \text{H} & -\text{C} & -\text{C} & =\text{C} & -\text{C} & -\text{C} & -\text{H} \\ & & & & & & \\ & \text{H} & & & & \text{H} & \end{array}$
	Name methylbut-2-ene (or 2-methylbut-2-ene)		

HALOGENOALKANES

	Structural formula	Skeletal formula	Name & comments
14	$\begin{array}{ccccccc} & & \text{CH}_3 & & & & \\ & & & & & & \\ & \text{CH}_3 & \text{CH}_2 & & \text{F} & & \\ & & & & & & \\ \text{CH}_3 - \text{CH} - \text{CH} - \text{CH}_2 - \text{CH} - \text{CH}_3 & & & & & & \end{array}$		<p>Longest C chain that contains the F functional group is 6C atoms – so it is has a <i>hex</i> stem.</p> <p>There is only one functional group (methyl and ethyl are not functional groups), and to give the F group the lowest number it is numbered from the right end, so it is <i>2-fluoro</i>.</p> <p>There is a methyl group on carbon-5, so it is <i>5-methyl</i>, and an ethyl group on carbon-4, so it is <i>4-ethyl</i>.</p> <p>In alphabetical order, it goes <i>4-ethyl</i> then <i>2-fluoro</i> and then <i>5-methyl</i>.</p> <p>4-ethyl-2-fluoro-5-methylhexane</p>
	Displayed formula		
	$\begin{array}{cccccccc} & & & & \text{H} & & & & \\ & & & & & & & & \\ & & & & \text{H} & & \text{H} & & \\ & & & & & & & & \\ & & & & \text{H} & & \text{H} & & \\ & & & & & & & & \\ \text{H} & - & \text{C} & - & \text{H} \\ & & & & & & & & & & & & & & \\ \text{H} & & \text{H} & & \text{H} & & \text{H} & & \text{F} & & \text{H} & & \text{H} & & \\ & & & & & & & & & & & & & & \\ \text{H} & & \end{array}$		

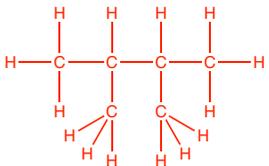
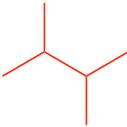
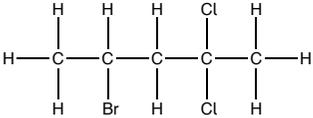
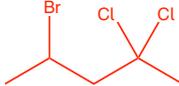
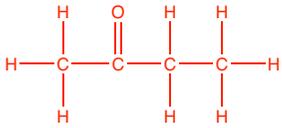
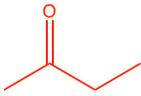
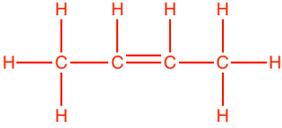
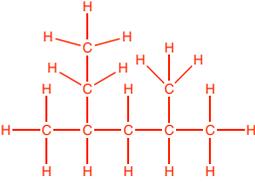
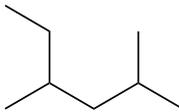
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15	$\text{CH}_3 - \text{CH}_2 - \text{Br}$		$\begin{array}{cc} \text{H} & \text{H} \\ & \\ \text{H} - \text{C} & - & \text{C} - \text{Br} \\ & \\ \text{H} & \text{H} \end{array}$
	Name bromoethane		

	Structural formula	Skeletal formula	Displayed formula
16	$\begin{array}{cccc} & \text{Cl} & & \\ & & & \\ \text{CH}_3 - \text{CH} & - & \text{C} & - \text{CH}_3 \\ & & & \\ & & & \end{array}$		$\begin{array}{cccc} \text{H} & \text{Cl} & & \text{H} \\ & & & \\ \text{H} - \text{C} & - & \text{C} & - & \text{C} & - & \text{C} & - & \text{H} \\ & & & \\ \text{H} & \text{H} & \text{I} & \text{H} \end{array}$
	Name 3-chloro-2,2-diiodobutane		

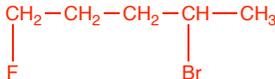
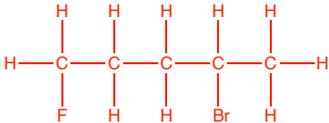
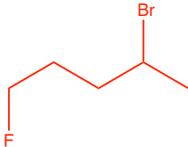
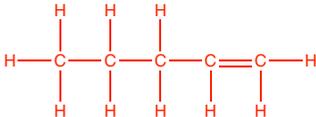
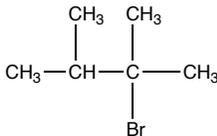
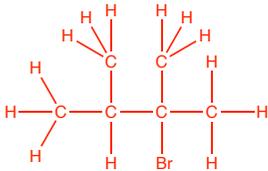
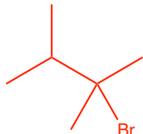
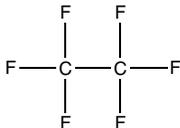
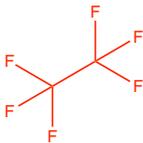
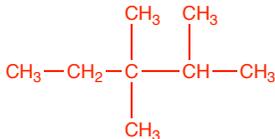
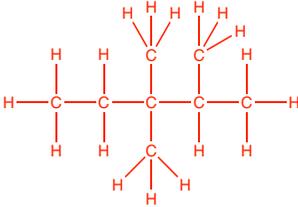
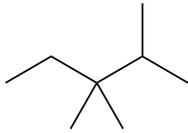
	Structural formula	Skeletal formula	Displayed formula
17	$\begin{array}{cccc} & \text{Br} & \text{Cl} & \text{CH}_3 \\ & & & \\ \text{CH}_3 - \text{C} & - & \text{CH} & - & \text{C} & - & \text{CH}_3 \\ & & & & \\ & \text{Br} & & \text{Br} & \end{array}$		$\begin{array}{ccccccc} & & & & \text{H} & & \\ & & & & & & \\ & & & & \text{H} & & \\ & & & & & & \\ & & & & \text{H} & & \\ & & & & & & \\ \text{H} & - & \text{C} & - & \text{H} \\ & & & & & & & & & & & & \\ \text{H} & & \text{Br} & & \text{Cl} & & \text{H} & & \text{Br} & & \text{H} & & \end{array}$
	Name 2,2,4-tribromo-3-chloro-4-methylpentane		

	Structural formula	Skeletal formula	Displayed formula
18	$\begin{array}{ccc} & \text{Cl} & \\ & & \\ \text{CH}_3 - \text{CH} & - & \text{CH} = \text{CH}_2 \end{array}$		$\begin{array}{cccc} \text{H} & \text{Cl} & \text{H} & \text{H} \\ & & & \\ \text{H} - \text{C} & - & \text{C} & = & \text{C} & - & \text{C} & - & \text{H} \\ & & & & & & & & \\ \text{H} & \text{H} & & & & & & & \end{array}$
	Name 3-chlorobut-1-ene		

TASK 1 – Nomenclature 1

Q	Name	Structural formula	Displayed formula	Skeletal formula
1.1	2,3-dimethylbutane	$ \begin{array}{cccc} \text{CH}_3 & -\text{CH} & -\text{CH} & -\text{CH}_3 \\ & & & \\ & \text{CH}_3 & \text{CH}_3 & \end{array} $		
1.2	4-bromo-2,2-dichloropentane	$ \begin{array}{ccccccc} & & & \text{Cl} & & & \\ & & & & & & \\ \text{CH}_3 & -\text{CH} & -\text{CH}_2 & -\text{C} & -\text{CH}_3 \\ & & & & \\ & \text{Br} & & \text{Cl} & \end{array} $		
1.3	butanone	$ \begin{array}{ccc} & \text{O} & \\ & & \\ \text{CH}_3 & -\text{C} & -\text{CH}_2-\text{CH}_3 \end{array} $		
1.4	but-2-ene	$ \text{CH}_3-\text{CH}=\text{CH}-\text{CH}_3 $		
1.5	2,4-dimethylhexane	$ \begin{array}{ccccccc} & & \text{CH}_3 & & & & \\ & & & & & & \\ & & \text{CH}_2 & & & & \\ & & & & \text{CH}_3 & & \\ \text{CH}_3 & -\text{CH} & -\text{CH}_2 & -\text{CH} & -\text{CH}_3 \\ & & & & & & \end{array} $		

1.6	methylpropene	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_2=\text{C}-\text{CH}_3 \end{array}$		
1.7	4-methylpent-2-ene	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{CH}-\text{CH}=\text{CH}-\text{CH}_3 \end{array}$		
1.8	1,1,2-tribromoethane	$\begin{array}{c} \text{Br} \\ \\ \text{CH}-\text{CH}_2 \\ \quad \\ \text{Br} \quad \text{Br} \end{array}$		
1.9	2,3-dimethylbut-1-ene	$\begin{array}{c} \text{CH}_3 \quad \text{CH}_3 \\ \quad \\ \text{CH}_2=\text{C}-\text{CH}-\text{CH}_3 \end{array}$		
1.10	2,2,4-trimethylpentane	$\begin{array}{c} \text{CH}_3 \quad \quad \text{CH}_3 \\ \quad \quad \\ \text{CH}_3-\text{CH}-\text{CH}_2-\text{C}-\text{CH}_3 \\ \\ \text{CH}_3 \end{array}$		

1.11	4-bromo-1-fluoropentane	$\text{CH}_2\text{FCH}_2\text{CH}_2\text{CHBrCH}_3$ 		
1.12	pent-1-ene	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}=\text{CH}_2$		
1.13	2-bromo-2,3-dibromobutane			
1.14	hexafluoroethane	CF_3-CF_3		
1.15	2,3,3-trimethylpentane			

ALCOHOLS

	Structural formula	Skeletal formula	Name & comments
19	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{CH}-\text{CH}-\text{CH}_3 \\ \\ \text{OH} \end{array}$		<p>Longest C chain that contains the OH functional group is 4C atoms – so it has a <i>but</i> stem.</p> <p>There is only one functional group (methyl is not a functional group), and to give the OH group the lowest number it is numbered from the right end, so it is <i>-2-ol</i>.</p> <p>There is a methyl group on carbon-3, so it is <i>3-methyl</i>.</p> <p>3-methylbutan-2-ol</p>
	Displayed formula		
	$\begin{array}{ccccccc} & & \text{H} & & \text{H} & & \\ & & & & & & \\ & & \text{C} & & \text{C} & & \\ & & / \backslash & & / \backslash & & \\ \text{H} & & \text{H} & & \text{H} & & \text{H} \\ & & & & & & \\ \text{H} & & \text{H} & & \text{O} & & \text{H} \\ & & & & & & \\ & & & & \text{H} & & \end{array}$		

	Structural formula	Skeletal formula	Name & comments
20	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{CH}-\text{CH}_2-\text{OH} \end{array}$		<p>Longest C chain that contains the OH functional group is 3C atoms – so it has a <i>prop</i> stem.</p> <p>There is only one functional group (methyl is not a functional group), and to give the OH group the lowest number it is numbered from the right end, so it is <i>-1-ol</i>.</p> <p>There is a methyl group on carbon-2. However, the methyl does not need a number in the name because it could not be on any other C atom and still be <i>methylpropan-1-ol</i>. However, many people would still put this 2 on the methyl group as it is not that obvious that a number is not needed</p> <p>methylpropan-1-ol (2-methylpropan-1-ol)</p>
	Displayed formula		
	$\begin{array}{ccccccc} & & \text{H} & & \text{H} & & \\ & & & & & & \\ & & \text{C} & & \text{C} & & \\ & & / \backslash & & / \backslash & & \\ \text{H} & & \text{H} & & \text{H} & & \text{O}-\text{H} \\ & & & & & & \\ \text{H} & & \text{H} & & \text{H} & & \end{array}$		

	Structural formula	Skeletal formula	Displayed formula
21	$\begin{array}{c} \text{CH}_3 \quad \text{CH}_3 \\ \quad \\ \text{CH}_3-\text{CH}-\text{CH}-\text{CH}-\text{CH}_2-\text{CH}_3 \\ \\ \text{OH} \end{array}$		
	Name 4,5-dimethylhexan-3-ol		

	Structural formula	Skeletal formula	Displayed formula
22	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_2-\text{CH}-\text{CH}=\text{CH}-\text{CH}_3 \\ \\ \text{OH} \end{array}$		
	Name 2-methylpent-3-en-1-ol		

	Structural formula	Skeletal formula	Displayed formula
23	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}-\text{OH} \\ \\ \text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}-\text{CH}_2-\text{CH}_3 \end{array}$		
	Name 3-ethylhexan-2-ol		

ALDEHYDES

24	Structural formula	Skeletal formula	Name & comments
	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{CH}-\text{CH}_2-\text{C}-\text{H} \\ \\ \text{O} \end{array}$		Longest C chain that contains the CHO aldehyde functional group is 4C atoms – so it is has a <i>but</i> stem.
	Displayed formula		By definition, the aldehyde functional group is carbon-1 and so it does not need a number. As the chain is numbered from the right, the methyl group is on carbon-3 and so it is <i>3-methyl</i> .
			3-methylbutanal

25	Structural formula	Skeletal formula	Displayed formula
	$\begin{array}{c} \text{O} \\ \\ \text{CH}_3-\text{CH}_2-\text{C}-\text{H} \end{array}$		
	Name propanal		

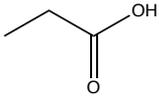
26	Structural formula	Skeletal formula	Displayed formula
	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{C}-\text{CH}-\text{CH}_2-\text{C}-\text{H} \\ \quad \\ \text{Br} \quad \text{Br} \\ \\ \text{O} \end{array}$		
	Name 3,4-dibromo-4-methylpentanal		

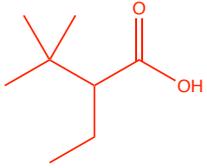
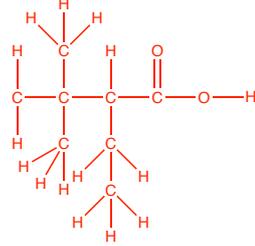
KETONES

27	Structural formula	Skeletal formula	Name & comments
	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{CH}-\text{CH}_2-\text{C}-\text{CH}_3 \\ \\ \text{O} \end{array}$		Longest C chain that contains the CO ketone functional group is 5C atoms – so it is has a <i>pent</i> stem.
	Displayed formula		There is only one functional group which is the ketone group (methyl is not a functional group). If the chain is numbered from the left, the ketone group is on carbon-4, but if it is numbered from the right it is carbon-2. Therefore we number from the right and it is <i>-2-one</i> . There is a methyl group on carbon-4 and so it is <i>4-methyl</i> .
			4-methylpentan-2-one

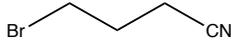
28	Structural formula	Skeletal formula	Displayed formula
	$\begin{array}{c} \text{O} \\ \\ \text{CH}_3-\text{CH}_2-\text{C}-\text{CH}_3 \end{array}$		
	Name butanone		

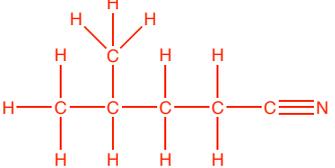
CARBOXYLIC ACIDS

	Structural formula	Skeletal formula	Name & comments
29	$\text{CH}_3-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$		<p>Longest C chain that contains the COOH carboxylic acid functional group is 3C atoms – so it has a <i>prop</i> stem.</p> <p>By definition, the carboxylic acid functional group is carbon-1 and so it does not need a number.</p>
	Displayed formula		
	$\begin{array}{c} \text{H} & \text{H} & \text{O} \\ & & \\ \text{H}-\text{C} & -\text{C}- & \text{C}-\text{O}-\text{H} \\ & & \\ \text{H} & \text{H} & \end{array}$	<p>propanoic acid</p>	

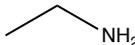
	Structural formula	Skeletal formula	Displayed formula
30	$\begin{array}{c} \text{CH}_3 & & \text{O} \\ & & \\ \text{CH}_3-\text{C} & -\text{CH}- & \text{C}-\text{OH} \\ & & \\ \text{CH}_3 & \text{CH}_2 & \\ & & \\ & \text{CH}_3 & \end{array}$		
	<p>Name 2-ethyl-3,3-dimethylbutanoic acid</p>		

NITRILES

	Structural formula	Skeletal formula	Name & comments
31	$\begin{array}{c} \text{Br} \\ \\ \text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CN} \end{array}$		<p>There are two functional groups, the –Br and the –CN.</p> <p>By definition, the nitrile functional group –CN is carbon-1 and so it does not need a number. This means that the Br is on carbon-4, so is <i>4-bromo</i>.</p>
	Displayed formula		
	$\begin{array}{c} \text{Br} & \text{H} & \text{H} \\ & & \\ \text{H}-\text{C} & -\text{C}- & \text{C}-\text{C}\equiv\text{N} \\ & & \\ \text{H} & \text{H} & \text{H} \end{array}$	<p>Note that the longest C chain containing the nitrile group has 4 C atoms, including the C of the nitrile group.</p> <p>4-bromobutanenitrile</p>	

	Structural formula	Skeletal formula	Displayed formula
32	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{CH}-\text{CH}_2-\text{CH}_2-\text{CN} \end{array}$		
	<p>Name 4-methylpentanenitrile</p>		

AMINES

	Structural formula	Skeletal formula	Name & comments
33	$\text{CH}_3-\text{CH}_2-\text{NH}_2$		<p>The naming of amines can be complicated and you are likely to see different names. The IUPAC rules for amines are not well known or used.</p> <p>This is a very simple amine with one alkyl group bonded to the N (a primary amine).</p> <p>The group bonded to the N is an ethyl group.</p> <p>It can be thought of as <i>ethane</i> but with an amine group in place of an H.</p>
	Displayed formula		
	$\begin{array}{c} \text{H} & \text{H} & \text{H} \\ & & \\ \text{H}-\text{C} & -\text{C}- & \text{N}-\text{H} \\ & & \\ \text{H} & \text{H} & \end{array}$	<p>ethanamine</p> <p><i>ethylamine or aminoethane also used</i></p>	

	Structural formula	Skeletal formula	Name & comments
34	$\begin{array}{c} \text{CH}_3-\text{CH}_2-\text{CH}-\text{NH}_2 \\ \\ \text{CH}_3 \end{array}$		<p>This molecule can be thought of as <i>butane</i> with an amine group on carbon-2. The amine group does need a number because it could also be on carbon-1.</p> <p>butan-2-amine</p> <p><i>1-methylpropylamine</i> or <i>2-aminobutane</i> also used</p>
	Displayed formula		

	Structural formula	Skeletal formula	Name & comments
35	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{NH}-\text{CH}_3$		<p>In this molecule, there are two alkyl groups bonded to the N and so it is a secondary amine.</p> <p>This molecule can be thought of as <i>propanamine</i> with a methyl group also bonded to the N. In order to show that it is bonded to the N rather than one of the C atoms in the propyl group, it is labelled <i>N-methyl</i>.</p> <p>N-methylpropan-1-amine</p> <p><i>N-methylpropylamine</i> also used</p>
	Displayed formula		

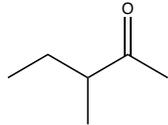
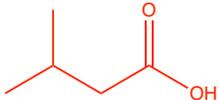
	Structural formula	Skeletal formula	Name & comments
36	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{N}-\text{CH}_3 \end{array}$		<p>In this molecule, there are three alkyl groups bonded to the N and so it is a tertiary amine.</p> <p>This is nearly always called <i>trimethylamine</i>, but it is not the official IUPAC name. It is effectively <i>methanamine</i> with two methyl groups on the N.</p> <p>N,N-dimethylmethanamine</p> <p><i>trimethylamine</i> usually used</p>
	Displayed formula		

	Structural formula	Skeletal formula	Displayed formula
37	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}-\text{NH}_2 \end{array}$		
	Name	2-methylbutanamine	

	Structural formula	Skeletal formula	Displayed formula
38	$\begin{array}{c} \text{CH}_3-\text{CH}-\text{CH}-\text{NH}-\text{CH}_3 \\ \quad \\ \text{CH}_3 \quad \text{CH}_3 \end{array}$		
	Name	N,2,3-trimethylpropanamine	

	Structural formula	Skeletal formula	Displayed formula
39	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{CH}_2-\text{CH}_2-\text{N}-\text{CH}_3 \end{array}$		
	Name	N,N-dimethylpropanamine (dimethylethylamine)	

TASK 2 – Nomenclature 2

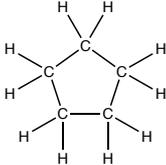
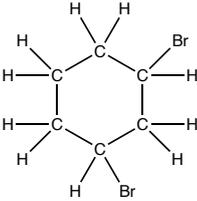
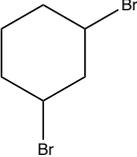
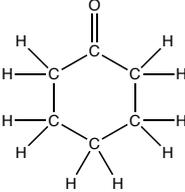
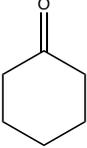
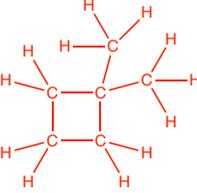
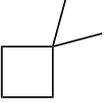
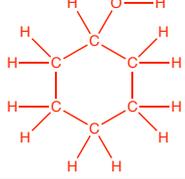
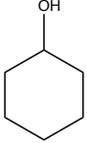
Q	Name	Structural formula	Displayed formula	Skeletal formula
2/1	3-methylpentan-2-one	$ \begin{array}{ccccccc} & & & \text{O} & & & \\ & & & & & & \\ \text{CH}_3 & - & \text{CH}_2 & - & \text{CH} & - & \text{C} & - & \text{CH}_3 \\ & & & & & & & & \\ & & & & \text{CH}_3 & & & & \end{array} $	$ \begin{array}{ccccccc} \text{H} & \text{H} & \text{H} & \text{O} & \text{H} & & \\ & & & & & & \\ \text{H}-\text{C} & - & \text{C} & - & \text{C} & - & \text{C} & - & \text{H} \\ & & & & & & & & \\ \text{H} & \text{H} & \text{H} & & \text{H} & & & & \\ & & & & & & & & \\ & & & & \text{C} & & & & \\ & & & & & & & & \\ & & & & \text{H} & & & & \\ & & & & & & & & \\ & & & & \text{H} & & & & \end{array} $	
2/2	3-methylbutan-1-ol	$ \begin{array}{ccccccc} \text{CH}_3 & - & \text{CH} & - & \text{CH}_2 & - & \text{CH}_2 \\ & & & & & & \\ & & \text{CH}_3 & & & & \text{OH} \end{array} $	$ \begin{array}{ccccccc} \text{H} & \text{H} & \text{H} & \text{H} & & & \\ & & & & & & \\ \text{H}-\text{C} & - & \text{C} & - & \text{C} & - & \text{C} & - & \text{H} \\ & & & & & & & & \\ \text{H} & \text{H} & \text{H} & \text{H} & & & \text{O} & - & \text{H} \\ & & & & & & & & \\ & & & \text{C} & & & & & \\ & & & & & & & & \\ & & & \text{H} & & & & & \\ & & & & & & & & \\ & & & \text{H} & & & & & \end{array} $	
2/3	ethanal	$ \begin{array}{ccc} & \text{O} & \\ & & \\ \text{CH}_3 & - & \text{C} & - & \text{H} \end{array} $	$ \begin{array}{ccc} \text{H} & \text{O} & \\ & & \\ \text{H}-\text{C} & - & \text{C} & - & \text{H} \\ & & & & \\ \text{H} & & & & \end{array} $	
2/4	3-methylbutanoic acid	$ \begin{array}{ccccccc} & & & \text{O} & & & \\ & & & & & & \\ \text{CH}_3 & - & \text{CH} & - & \text{CH}_2 & - & \text{C} & - & \text{OH} \\ & & & & & & & & \\ & & \text{CH}_3 & & & & & & \end{array} $	$ \begin{array}{ccccccc} \text{H} & \text{H} & \text{H} & \text{O} & & & \\ & & & & & & \\ \text{H}-\text{C} & - & \text{C} & - & \text{C} & - & \text{C} & - & \text{O} & - & \text{H} \\ & & & & & & & & & & \\ \text{H} & \text{H} & \text{H} & & \text{H} & & & & & & \\ & & & & & & & & & & \\ & & & & \text{C} & & & & & & \\ & & & & & & & & & & \\ & & & & \text{H} & & & & & & \\ & & & & & & & & & & \\ & & & & \text{H} & & & & & & \end{array} $	
2/5	butanenitrile	$ \text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CN} $	$ \begin{array}{ccccccc} \text{H} & \text{H} & \text{H} & & & & \\ & & & & & & \\ \text{H}-\text{C} & - & \text{C} & - & \text{C} & - & \text{C} \equiv \text{N} \\ & & & & & & \\ \text{H} & \text{H} & \text{H} & & & & \end{array} $	

2/6	propylamine or propanamine	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{NH}_2$		
2/7	2-methylpentan-3-one or methylpentan-3-one	$\text{CH}_3\text{CH}_2\text{COCH}(\text{CH}_3)_2$		
2/8	2-methylpropan-2-ol or methylpropan-2-ol			
2/9	dimethylamine or methylmethanamide	$\text{CH}_3-\text{NH}-\text{CH}_3$		
2/10	2,3-dimethylpentane			

2/16	butan-2-ol	$\begin{array}{c} \text{CH}_3 - \text{CH} - \text{CH}_2 - \text{CH}_3 \\ \\ \text{OH} \end{array}$		
2/17	hexanoic acid	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{COOH}$		
2/18	3-methylbutanenitrile	$\begin{array}{c} \text{CH}_3 - \text{CH} - \text{CH}_2 - \text{CN} \\ \\ \text{CH}_3 \end{array}$		
2/19	3-ethyl-2-methylpentanal	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_2 \\ \\ \text{CH}_3 - \text{CH}_2 - \text{CH} - \text{CH} - \text{C} = \text{O} \\ \quad \quad \quad \\ \text{CH}_3 \quad \quad \quad \text{H} \end{array}$		
2/20	2,4,4-trimethylpent-2-ene	$\begin{array}{c} \text{CH}_3 \\ \\ \text{H}_3\text{C} - \text{C} - \text{CH} = \text{C} - \text{CH}_3 \\ \quad \quad \\ \text{CH}_3 \quad \quad \text{CH}_3 \end{array}$		

CYCLIC COMPOUNDS

Note - it is normal to draw skeletal (or sometimes displayed) formulas for cyclic compounds – it is quite hard to draw anything else

40	Displayed formula	Skeletal formula	Name & comments
			<p>This is a cyclic alkane (there is no functional group). It has 5 C atoms in the ring.</p> <p>cyclopentane</p>
41	Displayed formula	Skeletal formula	Name & comments
			<p>This is a cyclic molecule with 6 C atoms in the ring, and so is <i>cyclohex</i>.</p> <p>There are two bromo functional groups. They could be on different C atoms on the ring and so numbers are needed. We take the first Br atom to be on carbon-1, and so the other Br atom is on carbon-3 going clockwise (or carbon-4 anticlockwise, but we go the way that gives the lowest number).</p> <p>1,3-dibromocyclohexane</p>
42	Displayed formula	Skeletal formula	Name & comments
			<p>This is a cyclic molecule with 6 C atoms in the ring, and so is <i>cyclohex</i>.</p> <p>There is a ketone functional group. It does not need a number as it is the only group on the ring.</p> <p>cyclohexanone</p>
43	Displayed formula	Skeletal formula	Name
			<p>1,1-dimethylcyclobutane</p>
44	Displayed formula	Skeletal formula	Name
			<p>cyclohexanol</p>

COMPOUNDS WITH SEVERAL FUNCTIONAL GROUPS

a) More than one of the same functional group

	Structural formula	Name
alkenes	$\text{CH}_3-\text{CH}=\text{CH}-\text{CH}_2-\text{CH}=\text{CH}_2$	hexa-1,4-diene
halogenoalkanes	$\begin{array}{c} \text{Br} \quad \text{Br} \\ \quad \\ \text{CH}_3-\text{CH}-\text{CH}-\text{CH}_2-\text{CH}_3 \end{array}$	2,3-dibromopentane
alcohols	$\begin{array}{c} \text{OH} \quad \text{OH} \\ \quad \\ \text{CH}_2-\text{CH}_2-\text{CH}_2 \end{array}$	propane-1,3-diol
aldehydes	$\begin{array}{c} \text{O} \quad \quad \quad \text{O} \\ \quad \quad \quad \\ \text{H}-\text{C}-\text{CH}_2-\text{CH}_2-\text{C}-\text{H} \end{array}$	butanedial
ketones	$\begin{array}{c} \text{O} \quad \quad \quad \text{O} \\ \quad \quad \quad \\ \text{CH}_3-\text{C}-\text{CH}_2-\text{C}-\text{CH}_3 \end{array}$	pentane-2,4-dione
carboxylic acids	$\begin{array}{c} \text{O} \quad \quad \quad \text{O} \\ \quad \quad \quad \\ \text{HO}-\text{C}-\text{CH}_2-\text{C}-\text{OH} \end{array}$	propanedioic acid
nitrile	$\text{NC}-\text{CH}_2-\text{CH}_2-\text{CN}$	butanedinitrile
amine	$\text{H}_2\text{N}-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{NH}_2$	propane-1,3-diamine
acyl chloride	$\begin{array}{c} \text{O} \quad \quad \quad \text{O} \\ \quad \quad \quad \\ \text{Cl}-\text{C}-\text{CH}_2-\text{C}-\text{Cl} \end{array}$	propanedioyl dichloride
amide	$\begin{array}{c} \text{O} \quad \quad \quad \text{O} \\ \quad \quad \quad \\ \text{H}_2\text{N}-\text{C}-\text{CH}_2-\text{C}-\text{NH}_2 \end{array}$	propanediamide

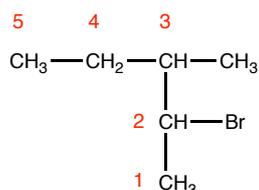
b) Different functional groups

- In terms of numbering the chain, there is an order of priorities with the highest priority groups taking the lowest number – you are not generally expected to know this order as an A level student, but it is useful to be aware of it – some common groups are shown here:
 carboxylic acid > nitrile > aldehyde > ketone > alcohol > amine > alkene > halogenoalkane
- If there are two groups that have suffixes, then the lower priority group becomes a prefix.

	Structural formula	Skeletal structure	Name & comments
45	$\begin{array}{c} \text{O} \quad \quad \quad \text{O} \\ \quad \quad \quad \\ \text{CH}_3-\text{C}-\text{CH}_2-\text{CH}_2-\text{C}-\text{H} \end{array}$		This has an aldehyde group and a ketone group. The aldehyde group takes priority and so is carbon-1. The ketone group is therefore carbon-4 and will be the prefix <i>4-oxo</i> 4-oxopentanal
46	$\begin{array}{c} \text{O} \quad \quad \quad \text{OH} \\ \quad \quad \quad \\ \text{CH}_3-\text{CH}_2-\text{C}-\text{CH}_2-\text{CH}-\text{CH}_3 \end{array}$		This has a ketone group and an alcohol group. The ketone group takes priority and so is carbon-3. The alcohol group is therefore carbon-5 and will be the prefix <i>5-hydroxy</i> 5-hydroxypentan-3-one

	Structural formula	Skeletal structure	Name & comments
47	$\text{CH}_3-\text{CH}=\text{CH}-\text{CH}_2-\underset{\text{OH}}{\text{CH}}-\text{CH}_3$		This has an alkene group and an alcohol group. The alcohol group takes priority and so is carbon-2 with suffix <i>-2-ol</i> . The alkene group is therefore carbon-4. hex-4-en-2-ol
48	$\text{H}_2\text{N}-\text{CH}_2-\text{CH}_2-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$		4-aminobutanoic acid
49	$\begin{array}{c} \text{Br} \quad \quad \quad \text{O} \\ \quad \quad \quad \parallel \\ \text{CH}_3-\text{CH}-\text{CH}-\text{C}-\text{CH}_3 \\ \\ \text{Br} \end{array}$		3,4-dibromopentan-2-one

Some more examples for extra guidance

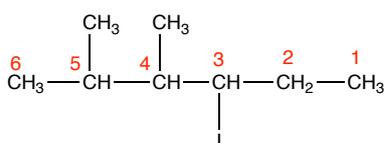


2-bromo-3-methylpentane

In this molecule, the longest C chain that contains the functional group has 5 C atoms (and so the name has *pent* in it.)

In order to give the functional group (bromo, **Br**) the lowest number, the C chain is numbered from the end shown. This makes it *2-bromo* rather than *4-bromo*.

There is also a *methyl* group on carbon-3 making it *3-methyl*. When writing the name, we put the *2-bromo* before the *3-methyl* the as *b* comes before *m* in the alphabet.



3-iodo-4,5-dimethylhexane

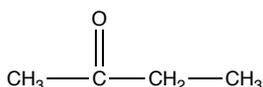
In this molecule, the longest C chain that contains the functional group has 6 C atoms (and so the name has *hex* in it.)

In order to give the functional group (iodo, **I**) the lowest number, the C chain is numbered from the right hand end. This makes it *3-iodo* rather than *4-iodo*.

There are also two *methyl* groups, on carbon-4 and carbon-5, making it *4,5-dimethyl*.

When writing the name, we put the *3-iodo* before the *4,5-dimethyl* the as *i* comes before *m* in the alphabet (we ignore the *di* in the dimethyl when doing this).

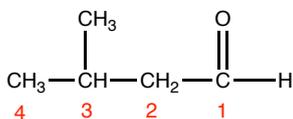
Note that the we would get a lower set of numbers if this was numbered from the other end, but we have to number the C chain to give the functional group the lowest number(s), and the only functional group here is the iodo group (methyl and any other alkyl groups are not functional groups).



butanone

In this molecule, the longest C chain that contains the functional group has 4 C atoms (and so the name has *but* in it.)

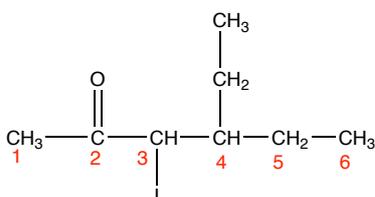
There is a ketone group (C=O) and so it has *-one* at the end. No number is needed for the ketone group because it can only be on one of the two central C atoms (otherwise it would be an aldehyde and not a ketone) and whichever of the two central C atoms it is, it would be carbon-2 and so no number is necessary.



3-methylbutanal

In this molecule, the longest C chain that contains the functional group has 4 C atoms (and so the name has *but* in it.)

The aldehyde group (CHO) is by definition carbon-1 and so the name has *al* at the end (note that no number is needed as the C in an aldehyde is on carbon-1 by definition). This means that the methyl group is on carbon-3 and is *3-methyl*.



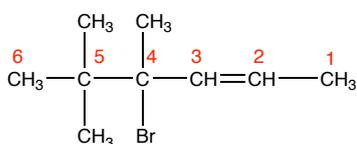
4-ethyl-3-iodohexan-2-one

In this molecule, the longest C chain that contains the functional group has 6 C atoms (and so the name has *hex* in it.)

There are two functional group (C=O and I). We number it from the end to give the ketone group the lowest number (as ketone groups have higher priority than halo groups). Therefore we number it from the left making it *2-one* and *3-iodo*.

There is also one *ethyl* group on carbon-4, making it *4-ethyl*.

When writing the name, we put the *4-ethyl* before the *3-iodo* the as e comes before i in the alphabet.



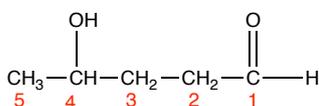
4-bromo-4,5,5-trimethylhex-2-ene

In this molecule, the longest C chain that contains the functional group has 6 C atoms (and so the name has *hex* in it.)

There are two functional group (alkene C=C and bromo Br). We number it from the end to give the alkene group the lowest number (as alkene groups have higher priority than halo groups). Therefore we number it from the right making it *2-ene* and *4-bromo*.

There are also three *methyl* groups, two on carbon-5 and one on carbon-4, making it *4,5,5-trimethyl*.

When writing the name, we put the *4-bromo* before the *4,5,5-trimethyl* the as b comes before m (ignore the tri) in the alphabet.



4-hydroxypentanal

In this molecule, the longest C chain that contains the functional group has 5 C atoms (and so the name has *pent* in it.)

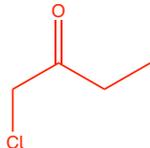
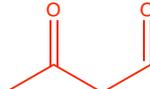
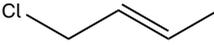
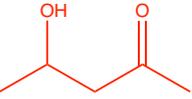
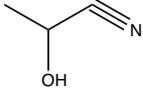
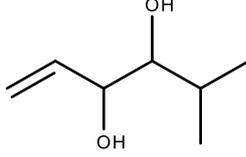
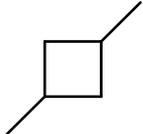
There are two functional groups (alcohol OH and aldehyde CHO). By definition, the aldehyde group is carbon-1 and so the alcohol group is carbon-4. (and aldehyde groups are higher priority than alcohol groups anyway)

Usually, these functional groups both are written as endings (ol and al). The aldehyde group is higher priority and so we use *al*, and so we use the prefix *4-hydroxy* for the alcohol.

TASK 3 – Nomenclature 3

Q	Name	Structural formula	Skeletal formula
3/1	4-hydroxypentanoic acid	$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3-\text{CH}-\text{CH}_2-\text{CH}_2-\text{C}-\text{OH} \\ \\ \text{OH} \end{array}$	
3/2	cyclohexanol		
3/3	but-3-enal	$\text{CH}_2=\text{CH}-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{H}$	
3/4	1,2-dibromocyclopentane		
3/5	4-amino-2-bromo-3-chlorohexan-3-ol	$\begin{array}{c} \text{NH}_2 \\ \\ \text{CH}_3-\text{CH}_2-\text{CH}-\text{C}-\text{CH}-\text{CH}_3 \\ \quad \quad \\ \text{OH} \quad \text{Cl} \quad \text{Br} \end{array}$	
3/6	2-fluorohex-3-ene	$\text{CH}_3\text{CH}_2\text{CH}=\text{CHCHFCH}_3$	
3/7	2,2-difluorobutanoic acid	$\begin{array}{c} \text{F} \quad \text{O} \\ \quad \parallel \\ \text{CH}_3-\text{CH}_2-\text{C}-\text{C}-\text{OH} \\ \\ \text{F} \end{array}$	
3/8	trichloroethanoic acid	CCl_3COOH	

3/9	but-3-en-2-ol	$\begin{array}{c} \text{CH}_2=\text{CH}-\text{CH}-\text{CH}_3 \\ \\ \text{OH} \end{array}$	
3/10	pentane-2,3-diol	$\begin{array}{c} \text{OH} \\ \\ \text{CH}_3-\text{CH}-\text{CH}-\text{CH}_2-\text{CH}_3 \\ \\ \text{OH} \end{array}$	
3/11	3-iodopentan-2-one	$\begin{array}{c} \text{I} \\ \\ \text{CH}_3-\text{C}-\text{CH}-\text{CH}_2-\text{CH}_3 \\ \\ \text{O} \end{array}$	
3/12	3-bromopropanal	$\text{CH}_2\text{BrCH}_2\text{CHO}$	
3/13	cyclopentene		
3/14	1,2-dibromocyclopropane		
3/15	hexa-1,3-diene	$\text{CH}_3-\text{CH}_2-\text{CH}=\text{CH}-\text{CH}=\text{CH}_2$	
3/16	hexane-1,3-dione	$\text{CH}_3-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_2-\text{CH}_3$	

3/17	pentanedinitrile	$\text{NC}-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CN}$	
3/18	1-chlorobutanone	$\begin{array}{c} \text{Cl} \\ \\ \text{CH}_2-\text{C}-\text{CH}_2-\text{CH}_3 \\ \\ \text{O} \end{array}$	
3/19	3-oxobutanal	$\begin{array}{c} \text{CH}_3-\text{C}-\text{CH}_2-\text{C}-\text{H} \\ \quad \quad \\ \text{O} \quad \quad \text{O} \end{array}$	
3/20	1-chlorobut-2-ene	$\begin{array}{c} \text{CH}_2-\text{CH}=\text{CH}-\text{CH}_3 \\ \\ \text{Cl} \end{array}$	
3/21	4-hydroxypent-2-one	$\begin{array}{c} \text{OH} \quad \quad \quad \text{O} \\ \quad \quad \quad \\ \text{CH}_3-\text{CH}-\text{CH}_2-\text{C}-\text{CH}_3 \end{array}$	
3/22	2-hydroxypropanenitrile	$\begin{array}{c} \text{CH}_3-\text{CH}-\text{C}\equiv\text{N} \\ \\ \text{OH} \end{array}$	
3/23	5-methylhex-1-ene-3,4-diol	$\begin{array}{c} \text{CH}_2=\text{CH}-\text{CH}-\text{CH}-\text{CH}-\text{CH}_3 \\ \quad \quad \\ \text{OH} \quad \text{OH} \quad \text{CH}_3 \end{array}$	
3/24	1,3-dimethylcyclobutane	$\begin{array}{c} \text{CH}_3 \\ \\ \text{H}_2\text{C}-\text{CH} \\ \quad \\ \text{HC}-\text{CH}_2 \\ \\ \text{H}_3\text{C} \end{array}$	

3/25	but-3-enoic acid	$\text{CH}_2=\text{CH}-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$	
3/26	cyclohexane-1,3-diol		
3/27	3-hydroxy-3-methylbutanoic acid	$(\text{CH}_3)_2\text{C}(\text{OH})\text{CH}_2\text{COOH}$	
3/28	5-fluorohexane-2,4-diol		
3/29	butanedioic acid		
3/30	ethane-1,2-diamine	$\text{H}_2\text{N}-\text{CH}_2-\text{CH}_2-\text{NH}_2$	
3/31	2-chloro-4-ethylhex-3-ene	$\text{CH}_3\text{CHClCH}=\text{C}(\text{C}_2\text{H}_5)_2$	
3/32	3-bromo-2-methyl-4-oxopentanal		

ACYL CHLORIDE

	Structural formula	Skeletal formula	Name & comments
50	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{Cl}$		The acyl chloride group is by definition on carbon-1 and so no number is needed. The longest chain is 4 C atoms long, so it is a <i>but.</i> butanoyl chloride
	Displayed formula		

	Structural formula	Skeletal formula	Displayed formula
51	$\text{CH}_3-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{Cl}$		
	Name propanoyl chloride		

ACID ANHYDRIDE

	Structural formula	Skeletal formula	Name & comments
52	$\text{CH}_3-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_2-\text{CH}_3$		Acid anhydrides are symmetrical about the middle oxygen (you could make acid anhydrides that are not, but there is no real use for them). The C chain length stem comes from the number of C atoms in the chain on each side of the central O. In this case, there are 3 C atoms on each side and so the stem is <i>prop.</i> propanoic anhydride
	Displayed formula		

	Structural formula	Skeletal formula	Displayed formula
53	$\text{CH}_3-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_3$		
	Name ethanoic anhydride		

ESTERS

	Structural formula	Skeletal formula	Name & comments
54	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}-\text{CH}_3$		There are two parts to the name. Esters can be thought of as being similar to carboxylic acids but with an alkyl group on the O instead of an H. This ester could be thought of as <i>butanoic acid</i> with a methyl group in place of the H on the O. The part of the molecule that is like <i>butanoic acid</i> without the H is <i>butanoate</i> . This <i>butanoate</i> has a <i>methyl</i> group attached. methyl butanoate
	Displayed formula		

	Structural formula	Skeletal formula	Displayed formula
55	$\text{CH}_3-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}-\text{CH}_2-\text{CH}_3$		
	Name ethyl propanoate		

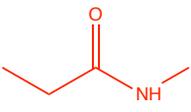
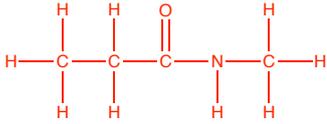
	Structural formula	Skeletal formula	Displayed formula
56	$\text{H}-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}-\text{CH}_2-\text{CH}_3$		
	Name ethyl methanoate		

AMIDES

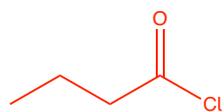
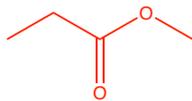
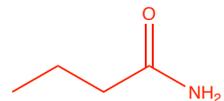
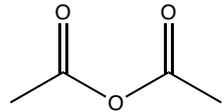
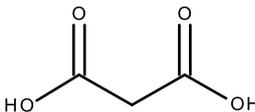
	Structural formula	Skeletal formula	Name & comments
57	$\text{CH}_3-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{NH}_2$		<p>The longest chain containing the amide functional group is 3 C atoms long, and so the stem is <i>prop</i>. The amide group is carbon-1 by definition and so no number is needed for it.</p> <p>propanamide</p>
	Displayed formula		

	Structural formula	Skeletal formula	Name & comments
58	$\text{CH}_3-\text{CH}_2-\underset{\text{CH}_3}{\text{CH}}-\overset{\text{O}}{\parallel}{\text{C}}-\text{NH}-\text{CH}_2-\text{CH}_3$		<p>The longest chain containing the amide functional group is 4 C atoms long, and so the stem is <i>but</i>. The amide group is carbon-1 by definition and so no number is needed for it.</p> <p>There is a methyl group on the main C chain. It could be on different C atoms and so needs a number. As the C of the amide group is carbon-1 by definition, the methyl group is on carbon-2 and so is <i>2-methyl</i>.</p> <p>There is an ethyl group on the N of the amide group. In order to show that this is on the N and not on a C atom, it is written as <i>N-ethyl</i>.</p> <p>In alphabetical order, the ethyl comes before the methyl.</p> <p>N-ethyl-2-methylbutanamide</p>
	Displayed formula		

	Structural formula	Skeletal formula	Displayed formula
59	$\text{CH}_3-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{NH}-\text{CH}_3$		
	Name		

	Structural formula	Skeletal formula	Displayed formula
60	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{NH}_2$		
	Name N-methyl propanamide		

TASK 4 – Nomenclature 4

Q	Name	Structural formula	Skeletal formula
4/1	butanoyl chloride	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{Cl}$	
4/2	methyl propanoate	$\text{CH}_3-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}-\text{CH}_3$	
4/3	butanamide	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{NH}_2$	
4/4	ethanoic anhydride	$\text{CH}_3-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_3$	
4/5	propanedioic acid	$\text{HO}-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$	
4/6	N-ethylpropanamide	$\text{CH}_3-\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{NH}-\text{CH}_2-\text{CH}_3$	